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Defect-engineered rGO–CoNi₂S₄ with enhanced electrochemical performance for asymmetric supercapacitor

Xu WANG¹, Bo-wang ZHAO¹, Jia-yu LIANG¹, Geng-zheng LIU¹,
Ze-fei GUO¹, Hui-lian HAO¹, Wen-yao LI¹, Wen-zhong SHEN²

1. College of Materials Materials Science and Engineering, Shanghai University of Engineering Science,
Shanghai 201620, China;

2. School of Physics and Astronomy, Shanghai Jiao Tong University, Shanghai 200240, China

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Abstract: To explore the effect of sulfur vacancies in transition metal sulfide on the electrochemical properties of anode materials, the graphene oxide (GO) and CoNi₂S₄ were used as the raw materials to synthesize the rGO₁₀–CoNi₂S_{4-x} composite electrode materials by the solvothermal method. The obtained rGO₁₀–CoNi₂S_{4-x} electrode materials with sulfur vacancies consist of nanoflakes and nanorods. The galvanostatic charge–discharge test of the rGO₁₀–CoNi₂S_{4-x} electrode materials shows a great specific capacitance of 3050.1 F/g at a current density of 1 A/g. Moreover, the electrode materials still remain rate capability retention of 86.1% when the current density increases from 1 to 10 A/g. The rGO₁₀–CoNi₂S_{4-x} composite containing sulfur vacancies has higher specific capacitance and better rate capability in comparison to the pristine rGO–CoNi₂S₄ without containing sulfur defects. The optimized rGO₁₀–CoNi₂S_{4-x} composite electrode materials with sulfur vacancies exhibit outstanding cycle stability and rate performance.

Key words: rGO₁₀–CoNi₂S_{4-x} composite; supercapacitor; electrochemical performance; sulfur vacancy; solvothermal method

1 Introduction

In recent years, considerable research has been conducted for searching the eco-friendly and high-efficiency electrochemical energy storage systems [1–3]. Supercapacitors (SCs), the novel form of energy storage devices with high power density, lying between batteries and conventional capacitors, have a potential application in the energy storage. However, SCs have a lower energy density, which severely hinders their widespread applications. Consequently, it is urgent to develop SCs with a high energy density to meet the requirements of energy storage devices. It is well known that the electrochemical performance of SCs relies heavily on the characteristics of electrode

materials, among which the transition metal sulfides have recently attracted much attention, due to their superior thermal stability, high theoretical capacitance, and exceptional electrical conductivity [3,4]. The nano-structured transition metal sulfides usually present prominent electrochemical properties. XU et al [5] synthesized partially amorphous CoNi₂S₄ nanosheets by a facile solvothermal method, which delivered a specific capacitance (C_s) of 3350 F/g at current density of 2 A/g. LIU et al [6] synthesized the uniformly reduced CoNi₂S₄ (r-CoNi₂S₄) nanosheets containing sulfur vacancies using a one-step hydrolysis process. The fabricated r-CoNi₂S₄ achieved C_s of 1918.9 F/g at 1 A/g and retained 76.2% of its original capacity after 10000 cycles. The formation of sulfur vacancies can significantly increase the number of sites and reduce the Gibbs

Corresponding author: Hui-lian HAO, Tel: +86-21-67791380, E-mail: sulee8866@126.com

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free energy of surface reactions, leading to the effective redox reactions [7,8]. However, transition metal sulfides are susceptible to agglomerate during charging–discharging processes, resulting in the reduction of surface area and active spots [9]. Thus, it is urgent to fabricate the electrode materials with outstanding electrochemical properties to satisfy the requirement of the high-performance supercapacitors. Defect engineering has been proven to be an efficient technique for creating active sites in transition metal sulfides and enhancing their intrinsic electrical conductivity [10,11]. The formation of sulfur vacancies in transition metal sulfides has been anticipated to increase the number of binding sites and conductivity, thus boosting the electrochemical properties [12,13].

Enhancing the conductivity and ion diffusion of transition metal sulfides can promote the charge transfer of electrode materials [14]. It is well known that the reduced graphene oxide (rGO) features with high conductivity and large specific surface area [9]. It is expected that the incorporation of low dimension and large specific surface area rGO into the electrode materials can further ameliorate the conductivity, ions diffusion rate and accessible surface areas of the faradic electrodes, which are essential factors to facilitate sufficient faradic reactions, therefore increasing the capacities of electrodes and devices. Additionally, the high dispersity of graphene oxide (GO) precursor can afford faradic composites, which can satisfy the comprehensive electrical performance demands in more widespread facilities.

In this work, the introduction of sulfur vacancies and rGO substantially enhances the faradic activity of the synthesized composite of rGO and CoNi_2S_4 (rGO– CoNi_2S_4) and restrains the agglomeration of transition metal sulfides during charge–discharge process to some degree. In addition, owing to the strong reduction characteristic of NaBH_4 , the obtained rGO– CoNi_2S_4 composite materials with sulfur vacancies preserve more reactive sites and low Gibbs free energy change, which can promote the electrochemical reactions and improve the electrochemical performance of the device. This research is expected to provide a theoretical and experimental basis for the research of transition metal sulfide electrode materials.

2 Experimental

2.1 Materials synthesis

The Ni–Co precursors were synthesized by a typical solvothermal method. In brief, 291.0 mg $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and 582.0 mg $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ were uniformly dissolved in the mixed solvent of ethylene glycol (EG) and deionized (DI) water (60 mL, $V_{\text{EG}}:V_{\text{DI}}=3:1$) by stirring for 30 min. Then, 378.2 mg $\text{C}_2\text{H}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ was slowly added into the above mixed solution under continuous stirring until it was completely dissolved to obtain the transparent mixture. The obtained mixture was subsequently transferred to an autoclave and treated hydrothermally for 24 h at 120 °C. Next, the reactants were cooled to 25 °C and cleaned with DI water and alcohol several times. The Ni–Co precursors were finally obtained by vacuum freeze-drying at –60 °C for 24 h.

GO was produced by the reported Hummers procedure. The rGO– CoNi_2S_4 composite with different GO masses was synthesized as follows: 300 mg Ni–Co precursor and a certain amount of GO were dispersed in the mixed solvent of EG (30 mL) and DI water (30 mL), and the uniform dispersion was achieved after 30 min ultrasonic treatment. Then, 300 mg $\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$ was added into the above dispersion and kept in a high-pressure reactor at 120 °C for 6 h to achieve the rGO– CoNi_2S_4 nanoflakes. The obtained composites were named as rGO_{*n*}– CoNi_2S_4 with different GO masses (*n*=0, 5, 10, 20 and 40 mg). The prepared rGO_{*n*}– CoNi_2S_4 samples were reduced by submerging in 0.9 mol/L NaBH_4 solution and placed at ambient temperature for 30 min to obtain rGO– CoNi_2S_4 with sulfur vacancies named as rGO_{*n*}– $\text{CoNi}_2\text{S}_{4-x}$ nanosheets. The preparation schematic diagram of rGO_{*n*}– $\text{CoNi}_2\text{S}_{4-x}$ composite material is shown in Fig. 1.

2.2 Material characterization

The structures of all the synthesized samples were analyzed using Raman spectroscopy (Raman, Horiba evolution), transmission electron microscopy (TEM, F200X), X-ray photoelectron spectroscopy (XPS, Thermo Scientific K-Alpha), scanning electron microscopy (SEM, ZEISS Gemini 300), X-ray diffraction (XRD, X’Pert PRO), electron

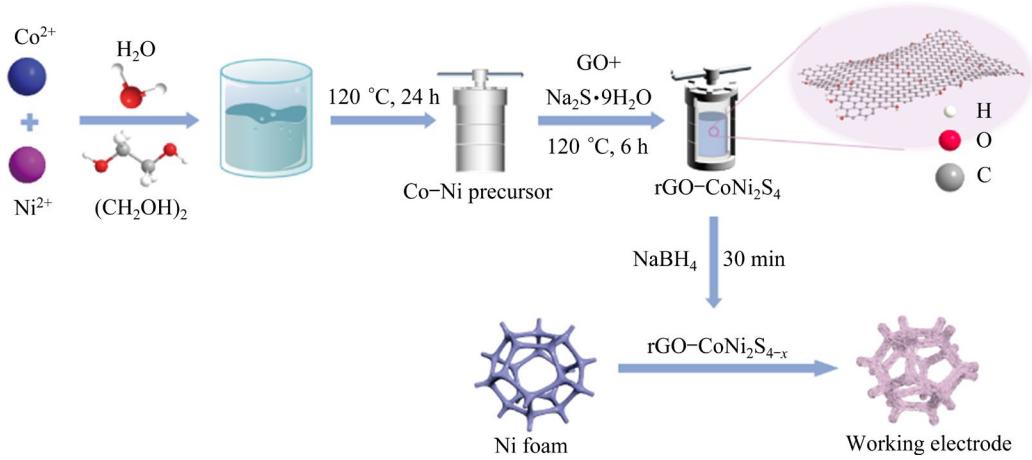


Fig. 1 Schematic illustration of synthesis of $\text{rGO}_n\text{--CoNi}_2\text{S}_{4-x}$

paramagnetic resonance (EPR, Bruker EMplus-6/1), and Brunauer–Emmet–Teller analyzer (BET, Micromeritics, JW-BK112T).

2.3 Electrochemical measurements

Electrochemical tests of all the as-prepared samples were performed using an electrochemical workstation (CHI660E, Shanghai, China) with a three-electrode setup. The three electrodes were working electrode, reference electrode and counter electrode, corresponding to the as-prepared electrode materials, Hg/HgO and platinum, respectively. Galvanostatic charge–discharge (GCD), electrochemical impedance spectroscopy (EIS), and cyclic voltammetry (CV) measurements were developed to assess the capabilities of electrode materials. The nickel foam (NF) was immersed in the mixture of HCl and ethanol solution to sonicate for 15 min, following vacuum-drying at 50 °C for 12 h. The pretreated NF was coated with the semi-solid slurry of the active material, acetylene black conductor, and polyvinylidene fluoride with a mass ratio of 8:1:1. Finally, the coated NF was dried at 60 °C for 12 h. Each electrode had about 1 cm × 1 cm loading area and the electrolyte was 3 mol/L KOH.

3 Results and discussion

3.1 Microstructure of electrode materials

SEM was used to observe the morphological structures of the electrode materials. Figure 2(a) displays the image of the Co–Ni precursors, which exhibits the interlaced linear structure with diameters of 100–200 nm. The micrograph of rGO in Fig. 2(b) displays an ultra-thin pleated structure.

The uniformly sized CoNi_2S_4 in Fig. 2(c) originates from the Co–Ni precursors after the sulfidation process, and the original geometry of Co–Ni precursors is well preserved. Figure 2(d) shows the SEM image of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ electrode, which exhibits a novel structure containing both nanoflakes and nanorods. The $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ presents smaller size and more irregular morphology, which is probably because the nuclei formation and crystal growth of the CoNi_2S_4 are affected by the groups of rGO during the reaction. It should be noted that the nanorods of CoNi_2S_4 are largely coated with rGO, which can serve as the protective layer to inhibit the severe volume variations during repeated electrochemical cycles, and therefore to improve the cycle ability of the electrode. Meanwhile, the surface of the $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ becomes rougher compared with that of CoNi_2S_4 , suggesting the larger specific surface area. The obtained $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ nanosheets still maintain partially open sheet structure, which is beneficial to $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ to shorten the ion diffusion distance and transport.

As determined by energy dispersive spectroscopy (EDS) mappings in Fig. 3, C, S, Ni, and Co components distribute uniformly throughout the $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ nanosheets. The presence of carbon element in EDS pattern of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ could be the further proof of the existence of rGO. In addition, EDS patterns in Fig. S1 (Supporting Information (SI)) clearly reveal that the Ni/Co:S molar ratio is significantly increased from 0.03 in CoNi_2S_4 to 0.13 in $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$, in accordance with the presence of an extensive number of sulfur vacancies [15].

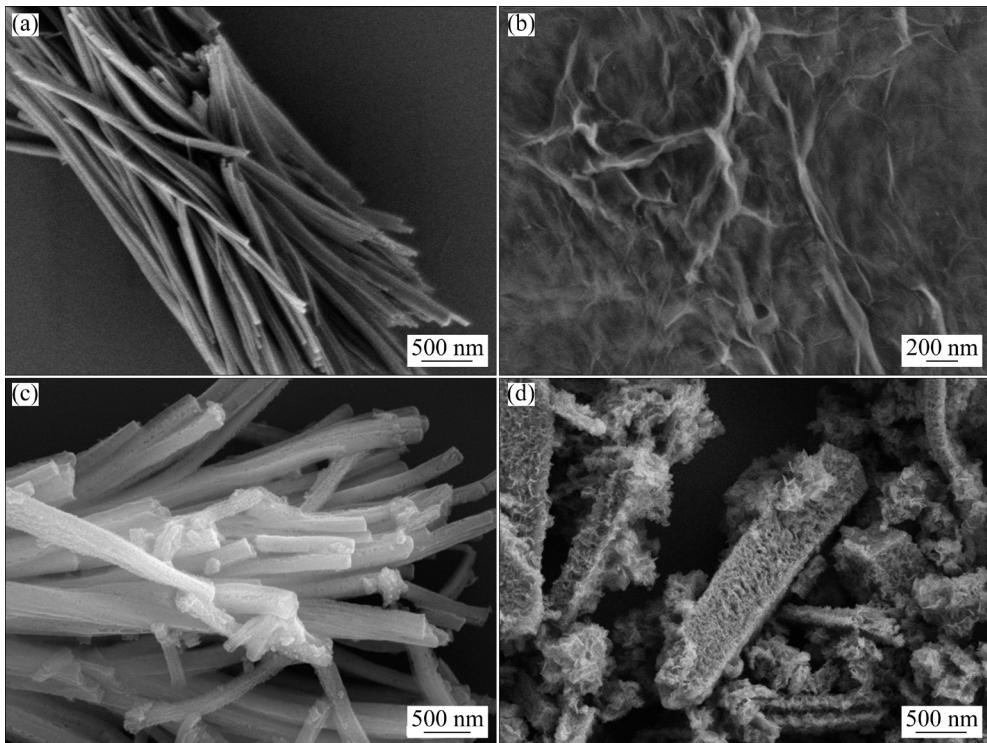


Fig. 2 SEM images of different materials: (a) Co–Ni precursor nanowires; (b) rGO; (c) Hierarchical CoNi₂S₄; (d) rGO₁₀–CoNi₂S_{4-x}

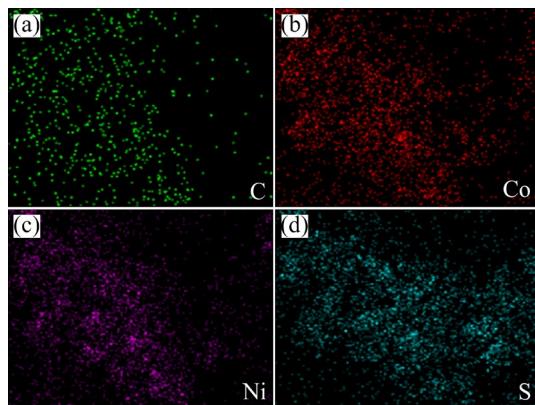


Fig. 3 EDS elemental mappings of rGO₁₀–CoNi₂S_{4-x}

Figure 4(a) displays the TEM image of rGO₁₀–CoNi₂S_{4-x}. It can be seen that rGO evenly develops on the surface of CoNi₂S_{4-x}, which can enhance the stability of the composite electrode material. The selected area electron diffraction (SAED) pattern of rGO₁₀–CoNi₂S_{4-x} in Fig. 4(b) shows the obvious multiple rings, indicating polycrystalline structure. The interlayer spacings (provided in Table S1 in SI) correlate well to the (440), (422), (400), and (220) of the composite. Through the high-resolution transmission electron microscopy (HRTEM, Fig. 4(c)) and inverse fast Fourier transform (IFFT, Figs. 4(c₁, c₂)) images, two groups of lattice stripes with interlayer spacings

of 0.28 and 0.33 nm correspond to (311) and (220) crystal planes of CoNi₂S₄, respectively. The above findings demonstrate the effective production of rGO₁₀–CoNi₂S_{4-x}.

Figure 5(a) demonstrates XRD diffraction peaks of CoNi₂S₄ and rGO₁₀–CoNi₂S_{4-x} at 2θ values of 26.7°, 31.5°, 38.0°, and 55.0°, which are corresponding to the (220), (311), (400), and (440) lattice planes of the cubic type CoNi₂S₄ (JCPDS No. 20-0334), respectively, good consistent with SAED results. Moreover, the well-matched XRD diffraction peaks for two samples indicate good preservation of the CoNi₂S₄ crystal structure following the treatment by NaBH₄. To detect the intrinsic defects of the sample, EPR measurements were performed to probe the sulfur vacancies (Fig. 5(b)). The gyromagnetic factor (*g*) governs the magnetic splitting of unpaired electron energy levels, with the free-electron theoretical value at about 2.0023. Deviations from this benchmark reflect localized chemical environments involving crystal field effects and spin-orbit coupling. Sulfur vacancies induce symmetry reduction through localized magnetic moments, directly modifying *g*-values via orbital reconfiguration. The reduced rGO₁₀–NiCo₂S_{4-x} sample displays an obvious signal at *g*=2.005, corresponding to the unpaired electrons

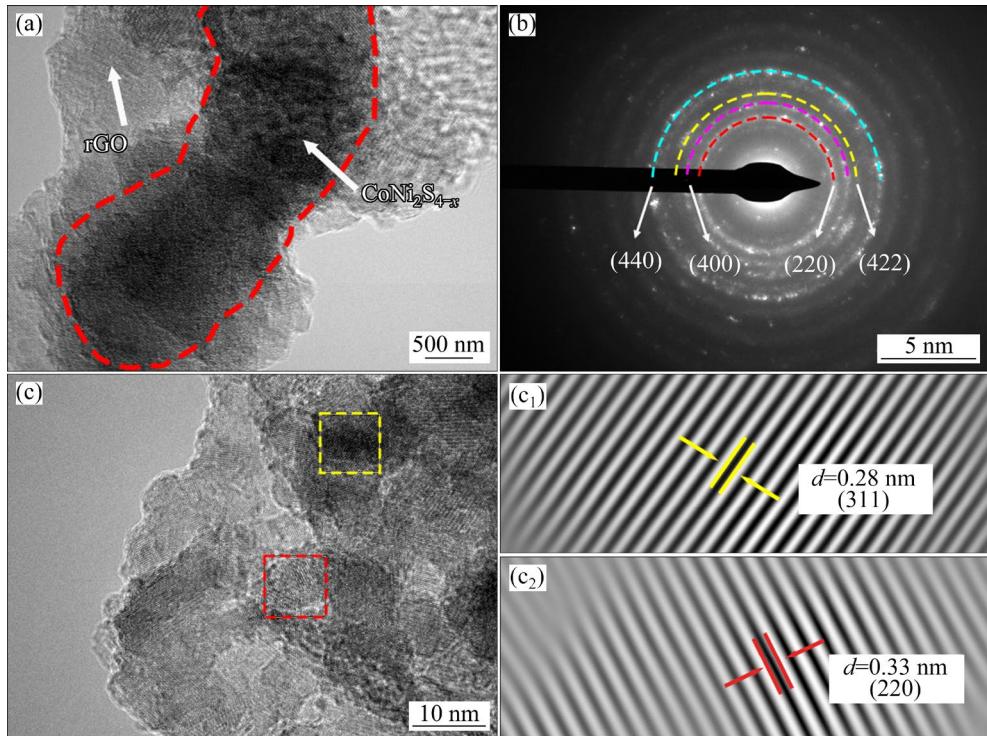


Fig. 4 Micro-morphologies of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$: (a) TEM image; (b) SAED pattern; (c) HRTEM image; (c₁, c₂) IFFT images

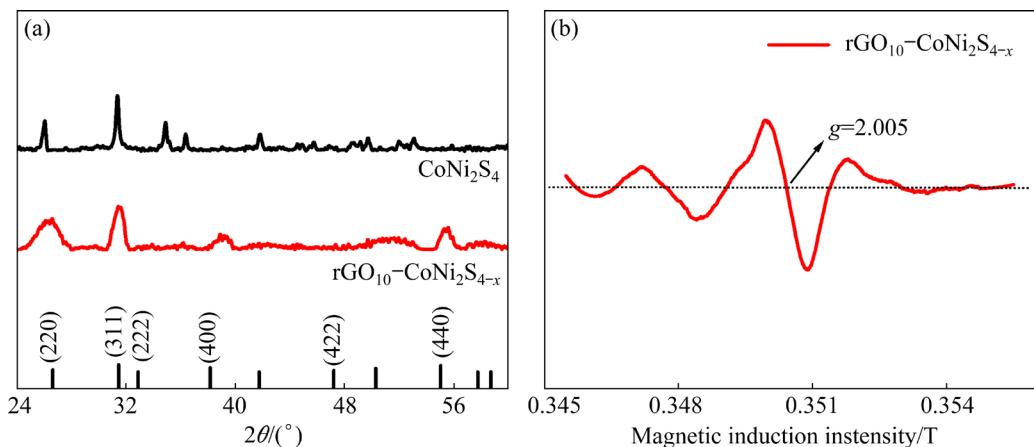


Fig. 5 XRD patterns of CoNi_2S_4 and $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ (a) and EPR pattern of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ (b)

trapped at the sulfur vacancies in the lattice [15], which also indicates that sulfur vacancies are created during solvothermal reduction. The presence of sulfur vacancies can significantly enhance redox rate and conductivity of composites, which is particularly beneficial to energy storage applications. Figure 6(a) presents the Raman spectra of the CoNi_2S_4 and $\text{rGO}_{10}\text{--NiCo}_2\text{S}_{4-x}$ nanosheets. For CoNi_2S_4 , the vibration peak at $\sim 464 \text{ cm}^{-1}$ is correlated to the E_g and A_g of S–S pairs in NiS_x , and the vibration peak at $\sim 515 \text{ cm}^{-1}$ is attributed to the S–S pair of CoS_x [16]. It can be obviously

found that the above two vibration peaks weaken or even disappear in $\text{rGO}_{10}\text{--NiCo}_2\text{S}_{4-x}$ compared with those of CoNi_2S_4 , demonstrating the reduction reaction of CoNi_2S_4 nanosheets and the formation of sulfur vacancies [17]. In addition, the peak located at $\sim 515 \text{ cm}^{-1}$ can be ascribed to the tetragonal S–Ni (Co) bands of NiCo_2S_4 and $\text{rGO}_{10}\text{--NiCo}_2\text{S}_{4-x}$, proving that the bimetallic sulfides have successfully grown onto the Ni foam [11]. Furthermore, the disappearance of vibration peak at $\sim 516 \text{ cm}^{-1}$ discloses the generation of sulfur vacancies [18]. As can be seen in Fig. 6(b), the

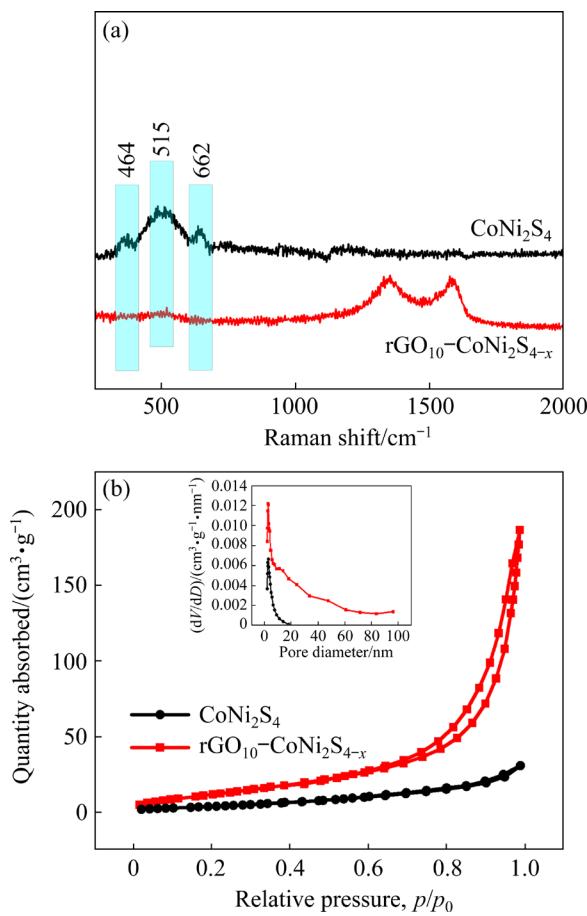


Fig. 6 Raman spectra (a) and N_2 adsorption–desorption isotherms (b) of $CoNi_2S_4$ and $rGO_{10}-CoNi_2S_{4-x}$

specific surface area and porous structure of electrode materials can be revealed by nitrogen adsorption–desorption experiments. It can be observed that both isotherms for $rGO_{10}-CoNi_2S_{4-x}$ and pure $CoNi_2S_4$ belong to IV type based on the Brunauer–Deming–Teller classification. The two obvious adsorption hysteretic loops locate at $p/p_0=0.7$ –1.0, suggesting that both samples owe porous structure with abundant mesopores. The calculated specific surface area of $rGO_{10}-CoNi_2S_{4-x}$ is $46.777\text{ m}^2/\text{g}$, which is much larger than that of $CoNi_2S_4$ ($15.275\text{ m}^2/\text{g}$). The inset of Fig. 6(b) illustrates the pore diameter (D) distribution according to the desorption isotherm, which is calculated based on Barret–Joyner–Halenda method. As can be seen from the inset, the sizes of mesopores lie 2–50 nm for both samples. The average pore diameter of $rGO_{10}-CoNi_2S_{4-x}$ ($\sim 24.7\text{ nm}$) is much larger than that of $CoNi_2S_4$ ($\sim 12.4\text{ nm}$). The above results suggest that the structure of $rGO_{10}-CoNi_2S_{4-x}$ is ideal for increasing active sites and conducive to the transport of

electrolyte ions, so as to achieve excellent electrochemical performance.

The elemental structure and chemical valence states of $CoNi_2S_4$ and $rGO_{10}-CoNi_2S_{4-x}$ nanosheet arrays were analyzed by XPS characterization. As can be seen from the full XPS spectra in Fig. 7, O, Co, Ni, S, and C elements can be detected for two samples. Figures 8(a–c) illustrate the typical XPS spectra of Co 2p, Ni 2p, and S 2p generated by Gaussian fitting for both samples, respectively. The full Co 2p spectra for two samples in Fig. 8(a) feature four spin-orbit locations and two satellite peaks (identified as “Sat.”). In the Co 2p spectrum of $CoNi_2S_4$, the two main peaks locate at ~ 781.3 and 797.4 eV with the spin-orbit splitting energy of 16.1 eV and correspond to Co^{2+} . Meanwhile, the two peaks situate at ~ 778.3 and 793.1 eV with the spin-orbit splitting energy around 14.8 eV and correlate to Co^{3+} [6].

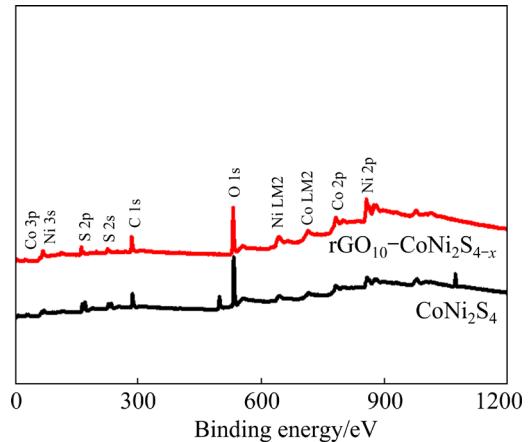


Fig. 7 XPS survey scan spectra for $CoNi_2S_4$ and $rGO_{10}-CoNi_2S_{4-x}$

Figure 8(b) shows the XPS spectra of Ni 2p for both samples. For $CoNi_2S_4$ sample, the two peaks at ~ 853.0 and 872.8 eV correspond to Ni^{2+} , and the double peaks at ~ 856.3 and 873.6 eV are related to Ni^{3+} . After $CoNi_2S_4$ sample is reduced by $NaBH_4$, the intensity of both peaks of Ni^{3+} in $rGO_{10}-CoNi_2S_{4-x}$ is much stronger than that of $CoNi_2S_4$, indicating the obvious increase of Ni^{3+} in $rGO_{10}-CoNi_2S_{4-x}$, which reveals that the content of Ni^{3+} is obviously increased. The increase of Ni^{3+} in $rGO_{10}-CoNi_2S_{4-x}$ can thereby facilitate the robust oxidation–reduction process and enhance the electrochemical properties of SCs [19]. For S 2p spectra of $CoNi_2S_4$ in Fig. 8(c), two peaks with binding energies at around 162.5 and 163.7 eV

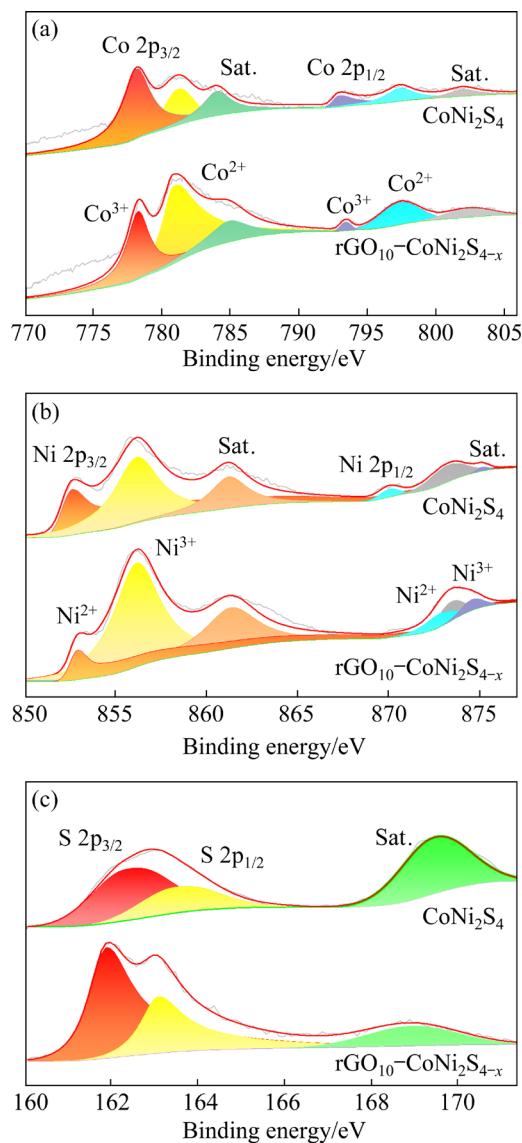


Fig. 8 XPS spectra of Co 2p (a), Ni 2p (b), and S 2p (c) of CoNi₂S₄ and rGO₁₀–CoNi₂S_{4-x} (Sat. means satellite peak)

correspond to S 2p_{3/2} and S 2p_{1/2}, respectively. In particular, the peak at 162.5 eV is ascribed to the metal sulfur bond in CoNi₂S₄ material [20]. Compared with the S 2p spectrum of CoNi₂S₄, the energy bands of S 2p_{3/2} and S 2p_{1/2} of rGO₁₀–CoNi₂S_{4-x} nanosheets take redshift to 161.9 and 163.1 eV, respectively, and the reduction in binding energies is attributed to the formation of sulfur vacancies [21]. According to first-principles analysis, the introduction of sulfur vacancies can considerably enhance the Fermi level of the S 2p states and charge transfer, therefore increasing the conductivity of rGO₁₀–CoNi₂S_{4-x} nanostructures [22].

3.2 Electrochemical properties of electrode materials

Figure 9(a) depicts the CV curves of the pristine Ni foam substrates, CoNi₂S₄ and rGO₁₀–CoNi₂S_{4-x} electrode materials at a scan rate of 10 mV/s in a range of 0–0.6 V. For bare Ni foam, the closed CV curve area is almost negligible compared with that of CoNi₂S₄ or rGO₁₀–CoNi₂S_{4-x}, illustrating that the capacitor contributor can be negligible for bare Ni foam and mainly originates from CoNi₂S₄ or rGO₁₀–CoNi₂S_{4-x}. For CoNi₂S₄ and rGO₁₀–CoNi₂S_{4-x} electrode materials, both CV curves show a pair of prominent symmetric redox peaks, the upper peaks are the oxidation peaks produced during the oxidation reaction and the lower ones are reduction peaks originating from the reduction reaction, demonstrating that the capacitance is predominant as a result of the reversible Faraday redox reaction. Due to the formation of sulfur vacancies, rGO₁₀–CoNi₂S_{4-x} composite material possesses a larger enclosed region of CV curve than CoNi₂S₄, revealing that much more charges can be stored. In addition, the redox peaks of rGO₁₀–CoNi₂S_{4-x} show slight shifts towards more positive and negative directions compared with those of CoNi₂S₄ under a specific scan rate, which can be attributed to the reversible redox reaction of nickel sulfide with the enhancement of Ni³⁺, as verified by the XPS measurements. The possible reaction based on Ni³⁺ can be excluded as follows: NiS+OH⁻ ⇌ NiSOH+e [22].

Figures 9(b) and S2(a) in SI illustrate the CV curves of rGO₁₀–CoNi₂S_{4-x} and CoNi₂S₄ electrode materials in the scan rate range of 1–30 mV/s, respectively. For all CV curves, the well-defined redox peaks originate from Faradic redox. With the enhancement of scan rate, on the one hand, the current intensity increases accordingly. On the other hand, it can be seen that the anodic peaks shift towards positive potential and the cathodic peaks move to negative potential. At lower scan rates, the redox reactions usually depend on the insertion–deinsertion of ions from the electrolyte, and the diffusion of ions from the electrolyte can enter into all the effective holes of the electrode. With increasing scan rate, the effective interaction between the ions and the electrode is greatly reduced, leading to a lower areal capacitance. The shapes of CV curves remain unchanged even at a

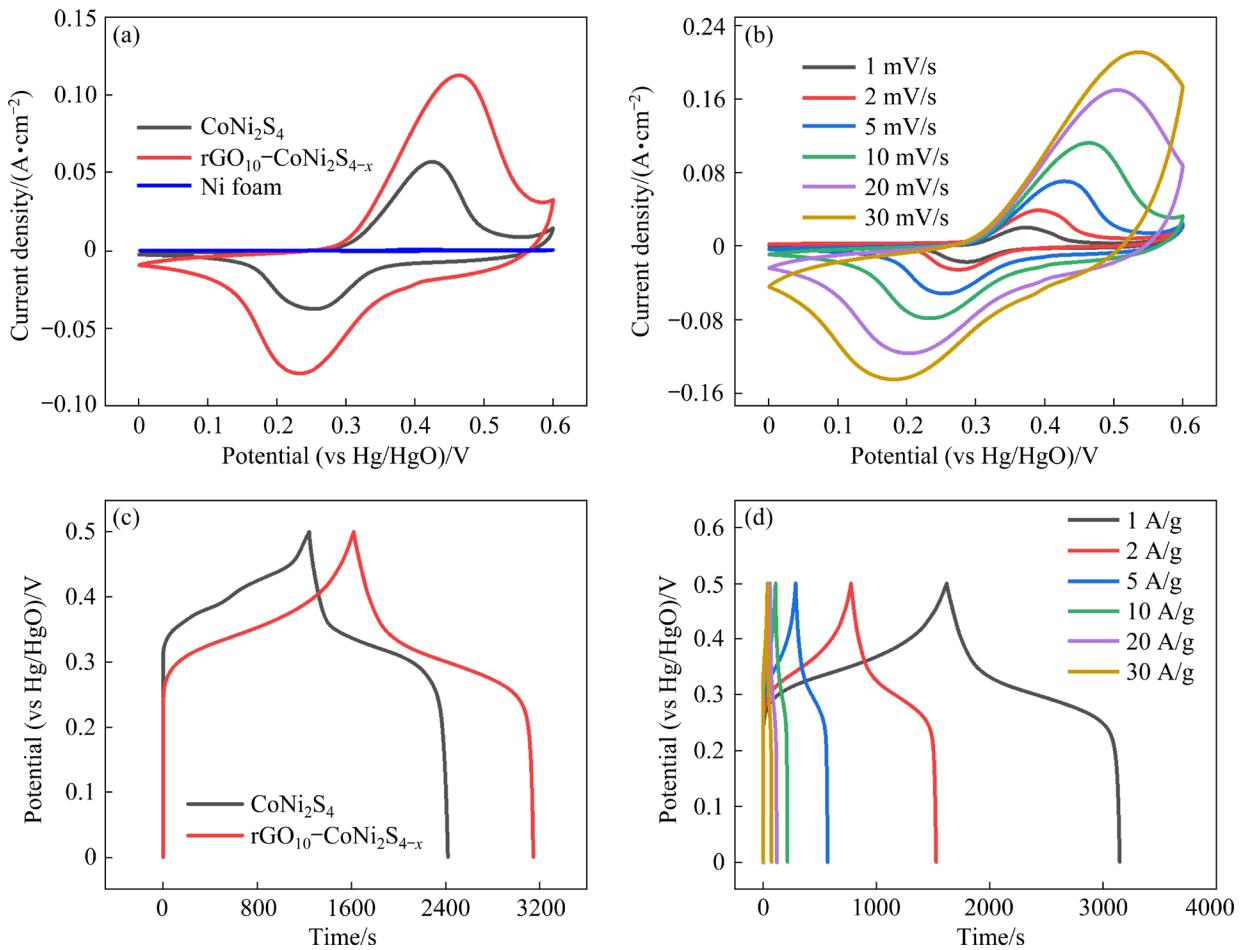
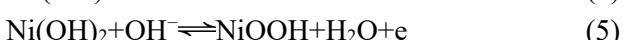
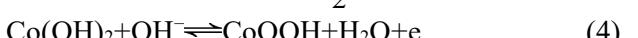
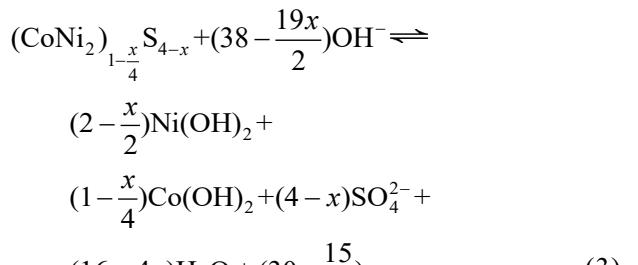
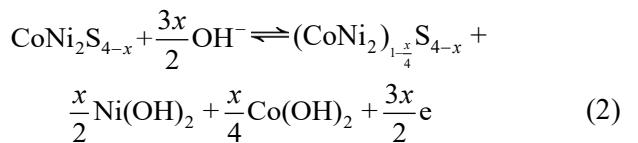


Fig. 9 CV curves of Ni foam, CoNi₂S₄ and rGO₁₀–CoNi₂S_{4-x} at scan rate of 10 mV/s in 3 mol/L KOH electrolyte (a), CV curves of rGO₁₀–CoNi₂S_{4-x} in scan rate range 1–30 mV/s (b), GCD curves of CoNi₂S₄ and rGO₁₀–CoNi₂S_{4-x} at 1 A/g (c), and GCD curves of rGO₁₀–CoNi₂S_{4-x} in current density range of 1–30 A/g (d)

scan rate as high as 30 mV/s, implying excellent reversibility and superior reaction kinetics. It should be noted that the closed CV curve area of rGO₁₀–CoNi₂S_{4-x} is much larger than that of CoNi₂S₄ at the corresponding scan rate, indicating much higher C_s . This can be attributed to the formed sulfur vacancies in rGO₁₀–CoNi₂S_{4-x} as proved by EPR measurements, which is conducive to promoting rapid redox capacity and favorable to be a candidate of electrode materials for supercapacitor [23]. The electrochemical reactions between rGO₁₀–CoNi₂S_{4-x} electrode materials and the electrolyte can be elucidated as follows [24]:



The intrinsic energy storage process of rGO₁₀–CoNi₂S_{4-x} in alkaline electrolytes primarily involves irreversible phase shifts, surface adsorption and partly reversible energy storage processes [6]:



GCD measurement is an effective method to study the energy storage capabilities of electrode materials. Figure 9(c) depicts the GCD test results for CoNi₂S₄ and rGO₁₀–CoNi₂S_{4-x} at 1 A/g. Due to the formation of sulfur vacancies as confirmed in

Figs. 5(b) and S1 in SI, $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ electrode has longer discharge time than CoNi_2S_4 electrode, indicating that more charges can be stored for the $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ electrode than CoNi_2S_4 , well consistent with CV results. It can be seen that for both samples, the obvious charging and discharging voltage platforms demonstrate the existence of the Faraday process, good correlation with CV results in Figs. 9(b) and Fig. S2(a) in SI. Moreover, the nearly symmetric charge–discharge profiles further demonstrate the highly reversible for the Faradaic redox reactions, as indicated by Reactions (2)–(5). Figures 9(d) and S2(b) in SI demonstrate the GCD curves of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ and CoNi_2S_4 throughout current densities of 1–30 A/g, respectively. All the GCD curves at different current densities remain almost the similar shapes for each sample, revealing the outstanding reversibility and rate performance. The discharge time of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ is longer than that of CoNi_2S_4 , indicating the efficient charge storage and transport. The C_s values shown in Fig. 10 were obtained from the charge–discharge profiles by Eq. (S1) in SI, and the corresponding C_s values of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ can reach 3050.1, 2999.3, 2809.7, 2625.5, 2368.3 and 2199.5 F/g at current densities of 1, 2, 5, 10, 20 and 30 A/g, respectively. These results are much higher than those of CoNi_2S_4 (2356.5, 1989.5, 1587.4, 1252.6, 868.0 and 557.5 F/g) under the corresponding current densities. Remarkably, with the current density increasing from 1 to 10 A/g, the specific capacitance of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ still maintains 86.1% at 10 A/g, which is higher than that of CoNi_2S_4 (53.2%), indicating the excellent rate performance. This can be explained as follows: on the one hand, the open sheet structure of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ in Fig. 2(d) can facilitate the fast transport of ions. On the other hand, the formation of sulfur vacancies in $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ composite electrode can provide the fast redox rate and good conductivity, thus enhancing the specific capacitance retention.

The long cycling life is very essential for supercapacitor applications. Figure 11(a) displays specific capacitance and coulombic efficiency obtained from the repeated GCD measurements at 10 A/g. The coulombic efficiency is also called the charge–discharge efficiency (η), which is defined as $\eta = (t_d/t_c) \times 100\%$, where t_d and t_c are the charge and discharge time, respectively. Following 10000

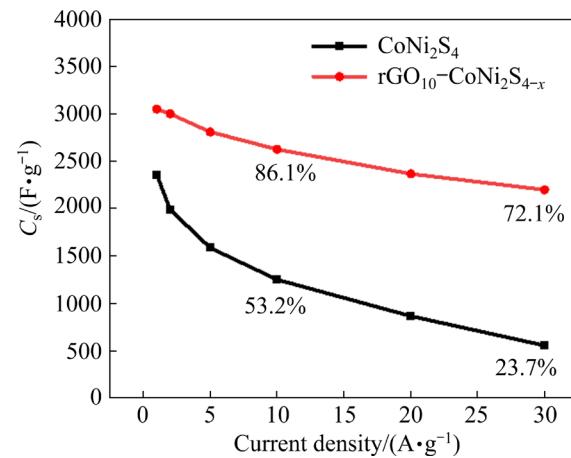


Fig. 10 C_s values of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ and CoNi_2S_4 at different current densities

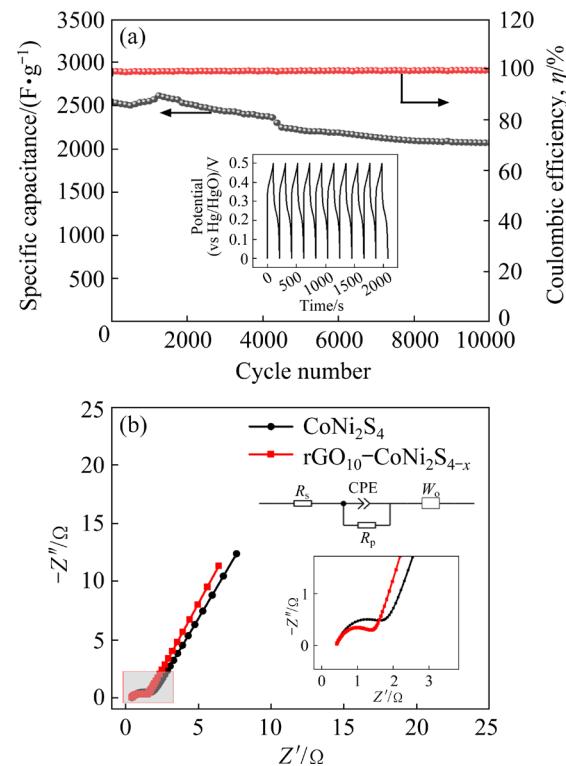


Fig. 11 Cycling stability of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ at 10 A/g for 10000 cycles (Inset: GCD curve of final 10 cycles) (a) and EIS curves of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ and CoNi_2S_4 (b)

charge–discharge cycles, the coulombic efficiency of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ keeps about 100%, and the capacitance retention still maintains 81.5% of its original value, indicating remarkable cycle stability and excellent electrochemical reversibility during the long-term cycle performance. The excellent electrochemical properties are of great significance for high-energy hybrid supercapacitors. Owing to the delayed activation during the cycles, the C_s of

$\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ increases a little bit for the first 2000 cycles, similar phenomena have been reported in Refs. [11,25]. The inset of Fig. 11(a) shows the charge–discharge curves for the last 10 cycles at 10 A/g, and the curve shape keeps almost unchanged, indicating good reversibility. For comparison, the η and cycling performance of CoNi_2S_4 are shown in Fig. S3 in SI. The calculated capacitance maintains almost 73.2% of its original value and coulombic efficiency is about 97.1% after 10000 cycles. The excellent performance of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ is closely correlated with the foamed nickel-supported porous sheet-like nanostructures of composite, which can greatly enhance the specific surface area and promote the transport of ions [26]. Meanwhile, the combination of graphene can effectively prevent the aggregation and volume expansion and contraction of composites during charging and discharging processes.

The charge transfer and ion diffusion mechanisms at the interface of electrolyte/electrode can be examined by EIS measurements. The impedance measurements of the $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ and CoNi_2S_4 nanosheets were performed in the frequency range of 0.01–10⁵ Hz. Figure 11(b) illustrates the Nyquist plots of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ and CoNi_2S_4 nano-composites together with the equivalent circuit, which comprises bulk solution resistance R_s , charge transfer resistance, constant phase element (CPE), polarization resistance R_p and Warburg impedance W_o . For two samples, both plots are composed of the semicircles and straight lines in the high and low frequency regions, respectively. The $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ electrode material has a lower R_s value of 0.404 Ω than CoNi_2S_4 (0.459 Ω), as obtained from the intercept of the Nyquist diagram with the real axis, suggesting that $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ electrode material owns fast electron transport kinetics at the interface of electrode and electrolyte [23]. Charge transfer resistance is represented by the semicircular diameter and corresponds to the interface resistance between active electrode materials and electrolyte. As can be seen from Fig. 11(b), $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ electrode material has a smaller semicircle compared to CoNi_2S_4 , demonstrating that charge transfer resistance value of the $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ (1.068 Ω) electrode material is lower than that of CoNi_2S_4 (1.398 Ω). The lower charge transfer resistance value represents the more effective

electrolyte/ion diffusion and faster charge transfer in the diffusion process [25]. The lower charge transfer resistance of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ electrode material can be attributed to the following reasons: firstly, rGO is well known to be a good electrical conductor, which can offer a low-resistive pathway for electrons, thus accelerating the reaction kinetics. Secondly, the existence of sulfur vacancies in $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ nano-composite is beneficial to the conductive contact between the nanosheets, as illustrated in Fig. 2(d) [13].

To gain an insight of electrochemical energy storage mechanism of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$, the diffusion-controlled and capacitance-controlled current contributions are investigated by the CV curves, as shown in Fig. 12(a). The current response can be described according to Eqs. (6) and (7) [27,28]:

$$I=av^b \quad (6)$$

$$\lg I=blg v+\lg a \quad (7)$$

where I (mA) is the sum of the diffusion-controlled current (I_{diff}) and the capacitance-controlled current (I_{cap}), v (mV/s) is the scan rate, and a and b mean adjustable parameters. The value of b is the slope of the linear fit of $\lg I$ vs $\lg v$. Typically, $b=0.5$ means the diffusion-controlled process, and $b=1$ suggests capacitive response. The slopes calculated from the linear fitting in Fig. S4 in SI related to the anode and cathode b values are 0.677 and 0.675, respectively. This indicates that the capacitance contributions originate from both the extrinsic capacitive and intrinsic diffusional charge–storage. The stored charge can be quantified based on Eq. (8) [29,30]:

$$I=I_{\text{cap}}+I_{\text{diff}}=k_1v+k_2v^{1/2} \quad (8)$$

where k_1v and $k_2v^{1/2}$ represent the capacitance and diffusion-controlled processes, respectively. k_1 and k_2 are the constants and can be obtained by plotting fitting lines of $I/v^{1/2}$ versus $v^{1/2}$, respectively. The capacitance contribution of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ at 10 mV/s is 58% (orange area). Meanwhile, the capacitive contributions for $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ at different scan rates are shown in Figs. 12(b) and S5 in SI, respectively. It can be obviously found that the capacitance contribution increases rapidly from 47% at the scan rate of 1 mV/s to 85% at 30 mV/s. The high capacitance contribution is essential to achieve the fast electrochemical kinetics for electrode

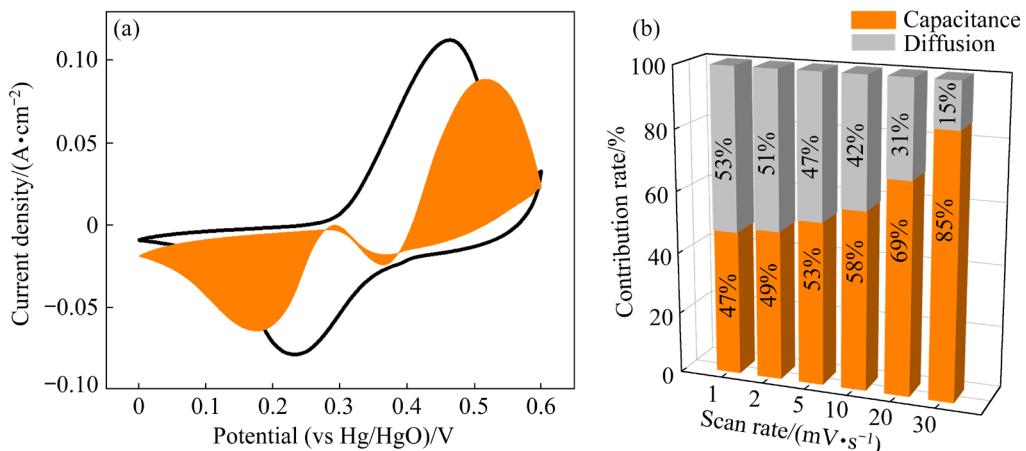


Fig. 12 Capacitance contribution of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ at scan rate of 10 mV/s (orange area) (a) and capacitance and diffusion distribution of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ at different scan rates

material in supercapacitors [31]. The larger capacitive contribution at higher scan rates can be attributed to the abundant porosity structure as illustrated in Fig. 6(b) and the introduction of sulfurs vacancies. Thus, high capacitance contributions are expected to improve electro-chemical performance and enhance stability.

3.3 Electrochemical performance of device

To further explore the realistic applications of the prepared $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ electrode materials on the supercapacitors, the $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}/\text{AC}$ hybrid ASC was built. Figure 13 schematically illustrates the structure and working mechanism of ASC, in which $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$, AC and 3 mol/L KOH are employed as positive, negative materials and electrolyte, respectively, one piece of cellulose paper is used as the separator. The mass ratio of the positive to negative electrodes is 1:11 based on the charge balance principle, as calculated by Eq. (S2) in SI.

The activated carbon is often employed as the anode material for energy storage devices, here the electrochemical properties of AC were examined with three-electrode work station. All the CV curves in Fig. S6(a) in SI present an ideal rectangular shape and keep almost unchanged in the scan rate range of 1–50 mV/s, indicating the rapid response behavior of ions and the capacitive characteristic. The GCD curves in Fig. S6(b) in SI at current densities of 1–10 A/g are approximately symmetric triangular, manifesting the typical capacitive performance. The obtained C_s values from GCD measurements in Fig. S6(c) in SI are

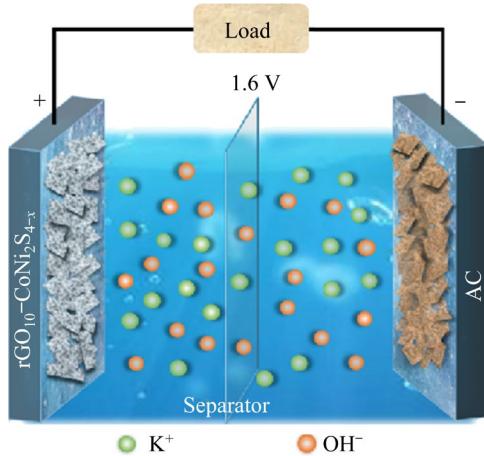


Fig. 13 Schematic illustration of assembled $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}/\text{AC}$ ASC

137, 129, 123, 118, 114 and 111 F/g at 1, 2, 4, 6, 8 and 10 A/g, respectively. Good conductivity can be revealed by smaller resistance shown in Fig. S6(d) in SI with smaller resistance. Thus, AC is expected to be an ideal anode material for SCs.

Figure 14 depicts CV curves of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}$ and AC electrodes with the corresponding working voltage of 0–0.6 V and -1–0 V, respectively. As a result, 1.6 V is expected to be the stable potential window for $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}/\text{AC}$ ASC. The CV curves in Fig. 15(a) of $\text{rGO}_{10}\text{-CoNi}_2\text{S}_{4-x}/\text{AC}$ device were obtained under different potential windows within 1.2–1.9 V to evaluate the optimal working voltage. It can be obviously seen that all the CV curves keep the ideal shape under the desired potential window of 1.6 V. The curves become significantly polarized when the voltage exceeds 1.7 V, which may result from the oxygen evolution reaction of the aqueous electrolyte [32].

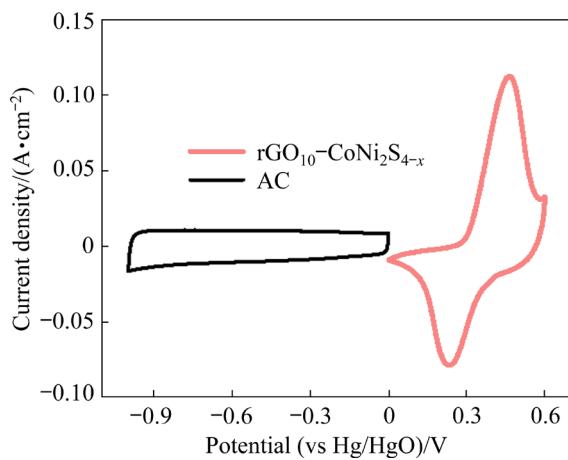


Fig. 14 CV curves of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ and AC electrodes at scan rate of 10 mV/s

Thus, the electrochemical properties of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}\text{/}/\text{AC}$ ASC were further characterized at the achieved voltage 1.6 V. Figure 15(b) provides the CV curves of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}\text{/}/\text{AC}$ ASC in the scan speed range of 1–50 mV/s, which reveals a quasi-rectangular shape with the prominent redox

peaks. All the CV curves keep almost the same shape at different scan rates, demonstrating invertible and fast charge and discharge characteristics. Figure 15(c) illustrates the GCD curves detected at different current densities. The quasi-symmetrical shape manifests the capacitive behavior originating from $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}$ nanocomposite, good consistent with the results shown in Fig. 12(b). The obtained C_s of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}\text{/}/\text{AC}$ ASC is 226 F/g at 1 A/g, and it maintains a capacitance retention of 60.6% at 10 A/g (137 F/g at 10 A/g) as in Fig. 15(d), implying excellent reversibility. Figure 16(a) shows the Nyquist diagram of the ASC within a frequency range from 10^{-2} to 10^5 Hz. To obtain EIS data, an equivalent circuit is established (the inset) and the calculated R_s value is 1.5 Ω , reflecting the low solution resistance.

Figure 16(b) illustrates the cycling performance of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}\text{/}/\text{AC}$ ASC. The capacitance retention reaches as high as 86.6% even after 10000 consecutive charge–discharge cycles at a current

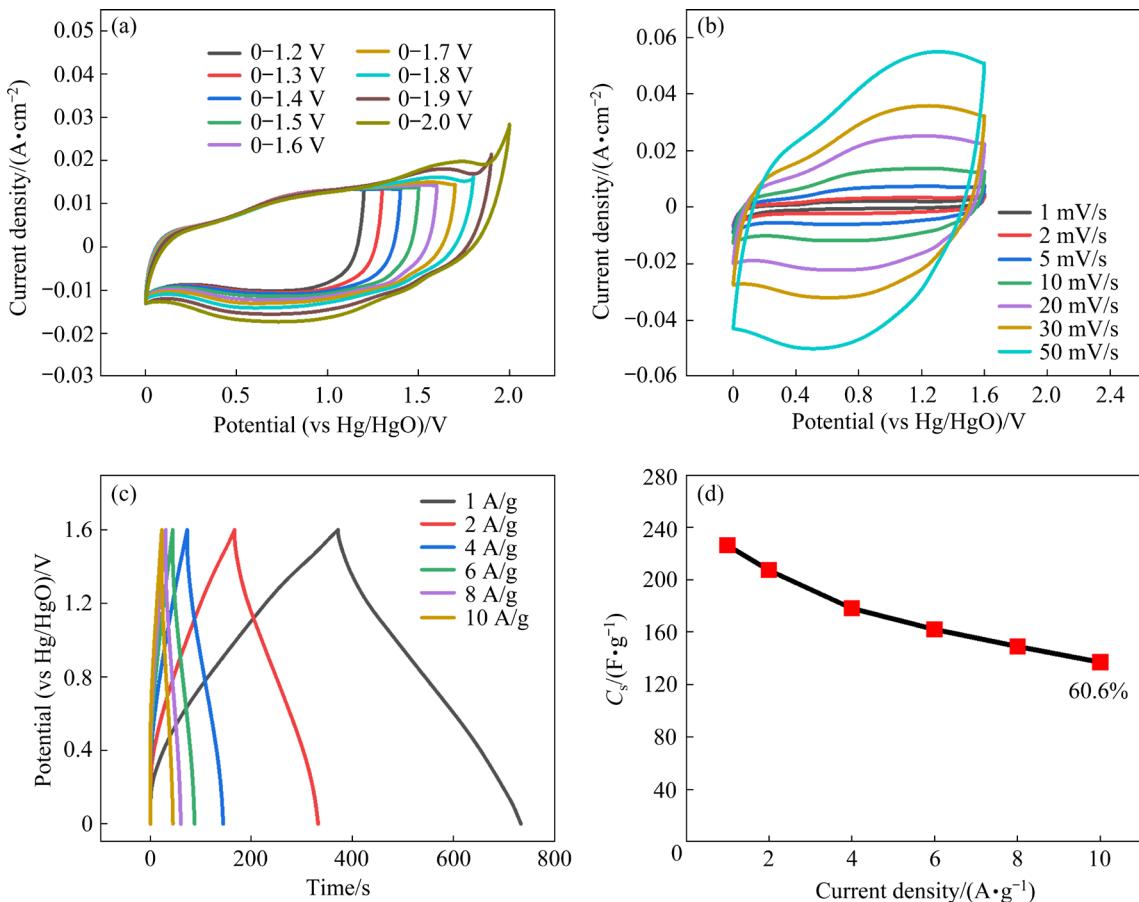


Fig. 15 CV curves of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}\text{/}/\text{AC}$ ASC at various potential windows and 10 mV/s (a), CV curves at 1–50 mV/s (b), GCD curves of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}\text{/}/\text{AC}$ ASC in current intensity range of 1–10 A/g (c), and C_s values of $\text{rGO}_{10}\text{--CoNi}_2\text{S}_{4-x}\text{/}/\text{AC}$ ASC at different current densities (d)

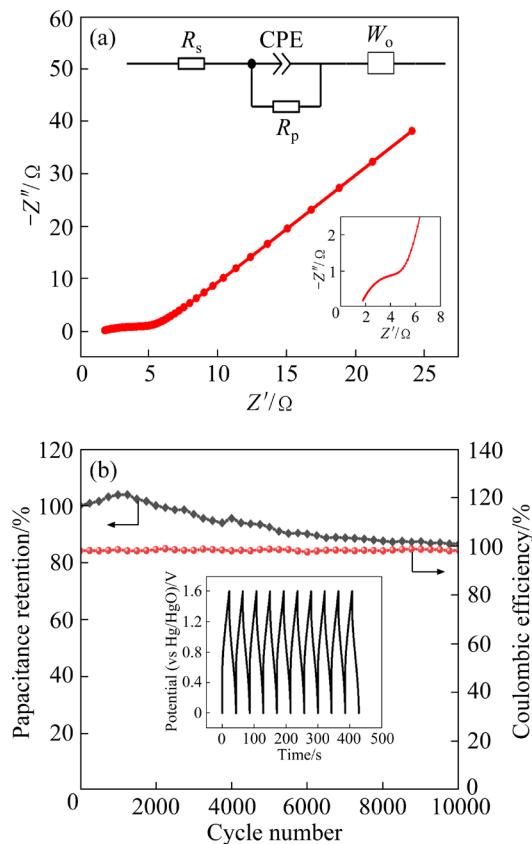


Fig. 16 EIS curve (a) and cycling performance of last 10 cycles (inset) at current density of 10 A/g (b) for $\text{rGO}_{10}-\text{CoNi}_2\text{S}_{4-x}/\text{AC}$ ASC

density of 10 A/g, indicating the excellent cycling stability. However, the de-sulphuration reaction between cobalt nickel sulfide and OH^- will cause the irreversible capacitance, resulting in the slight decrease of the capacitance retention. In the first 2500 cycles, due to the slow activation of the electrode, the capacitance rises a little bit and then falls until the steady state. Similar phenomena were reported in Refs. [25,33]. At the same time, it can also be found that coulombic efficiency always keeps 100% over the entire measurements, demonstrating the outstanding reversibility of the constructed device. This can be ascribed to the sufficient electrolyte penetration from the stacked nanosheet structure in the cycle process.

It is well known that both the energy density and power density are the very key parameters to evaluate the performance of energy storage devices. The energy density and power density of ASC can be calculated from Eqs. (S3) and (S4) in SI, as displayed by the Ragone diagrams in Fig. 17. The $\text{rGO}_{10}-\text{CoNi}_2\text{S}_{4-x}/\text{AC}$ ASC achieves a high energy

density of 80.35 W·h/kg at a power density 800 W/kg, and it still exhibits a high energy density up to 52.98 W·h/kg even at a power density 6400 W/kg. In this work, the as-constructed $\text{rGO}_{10}-\text{CoNi}_2\text{S}_{4-x}/\text{AC}$ ASC presents the much higher energy density and power density than the reported CoNiS-based asymmetric supercapacitors [21,34–37]. For a clearer comparison, Table S2 in SI lists the present work and reported energy storage properties of cobalt-nickel sulfides. The excellent electrochemical performance of $\text{rGO}_{10}-\text{CoNi}_2\text{S}_{4-x}/\text{AC}$ ASC mainly stems from the porous nanosheet cathode materials, which can supply sufficient charge transfer paths to promote Faraday redox reactions [38]. Meanwhile, the open nanostructure and the formed sulfur vacancies in $\text{rGO}_{10}-\text{CoNi}_2\text{S}_{4-x}$ can store much more charge [39]. Furthermore, the as-construed $\text{rGO}_{10}-\text{CoNi}_2\text{S}_{4-x}/\text{AC}$ ASC can easily drive a bulb at a working voltage of 1.5 V, as shown in Fig. S7 in SI, suggesting its practical applications to the supercapacitor.

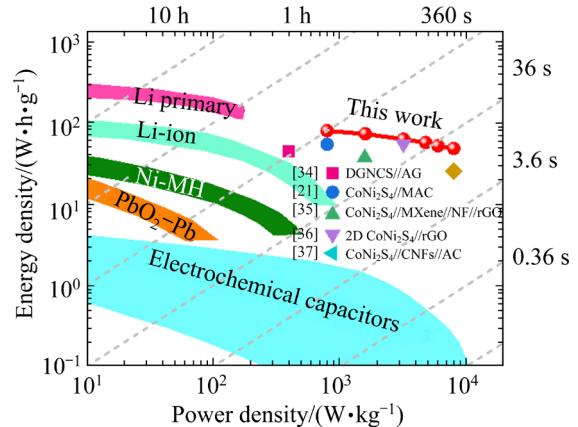


Fig. 17 Ragone plot of this work together with reported results

4 Conclusions

(1) The porous nanosheet $\text{rGO}_n-\text{CoNi}_2\text{S}_{4-x}$ composites with sulfur vacancies were synthesized through two-step hydrothermal procedure and a single-step reduction reaction.

(2) The formation of sulfur vacancies in $\text{rGO}_n-\text{CoNi}_2\text{S}_{4-x}$ composite can be confirmed by ESR results and Raman spectra.

(3) The fabricated $\text{rGO}_{10}-\text{CoNi}_2\text{S}_{4-x}/\text{AC}$ ASC delivers a high energy density of 80.35 W·h/kg at a power density of 800 W/kg and excellent

cycling stability (86.6%) ever after 10000 consecutive charge–discharge cycles at a high current density of 10 A/g.

CRediT authorship contribution statement

Xu WANG: Conceptualization, Methodology, Investigation, Structural and electrochemical measurements, Writing – Original draft; **Bo-wang ZHAO:** SEM and TEM measurements, Writing – Review & editing; **Jia-yu LIANG:** XRD and Raman measurements, Writing – Review & editing; **Geng-zheng LIU:** XPS measurements, Writing – Review & editing; **Ze-fei GUO:** TEM measurements, Writing – Review & editing; **Hui-lian HAO:** Resources, Writing – Review & editing, Supervision, Data curation; **Wen-yao LI:** Resources, Writing – Review & editing; **Wen-zhong SHEN:** Supervision, Data curation.

Declaration of competing interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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Supporting Information

Supporting information in this paper can be found at: http://tnmsc.csu.edu.cn/download/17-p0563-2023-0747-Supporting_Information.pdf.

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基于 rGO–CoNi₂S₄ 缺陷工程增强的 非对称超级电容器的电化学性能

王 翳¹, 赵伯旺¹, 梁佳宇¹, 刘庚正¹, 郭泽飞¹, 郝惠莲¹, 李文尧¹, 沈文忠²

1. 上海工程技术大学 材料科学与工程学院, 上海 201620;
2. 上海交通大学 物理与天文学院, 上海 200240

摘要: 为了探索过渡金属硫化物中的硫缺陷对其负极材料电化学性能的影响, 以氧化石墨烯(GO)和钴镍硫化物(CoNi₂S₄)为原料通过溶剂热法合成了含硫缺陷的还原氧化石墨烯–钴镍硫化物(rGO₁₀–CoNi₂S_{4-x})复合电极材料。所合成的 rGO₁₀–CoNi₂S_{4-x} 复合电极材料含有纳米片和纳米棒。恒电流充放电测试表明, 当电流密度为 1 A/g 时, rGO₁₀–CoNi₂S_{4-x} 电极材料的比电容为 3050.1 F/g; 当电流密度从 1 A/g 增加至 10 A/g 时, 倍率容量保持率为 86.1%。与不含硫缺陷的 rGO–CoNi₂S₄ 相比, 含有硫缺陷的 rGO₁₀–CoNi₂S_{4-x} 具有较高的比电容和较好的倍率性能。优化后的含硫缺陷 rGO₁₀–CoNi₂S_{4-x} 电极材料呈现良好的循环性能和倍率性能。

关键词: rGO₁₀–CoNi₂S_{4-x} 复合材料; 超级电容器; 电化学性能; 硫缺陷; 溶剂热法

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