

Kinetics analysis of non-isothermal decomposition of Ag_2O –graphite mixture

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Abstract: The thermal decomposition kinetics of Ag_2O –graphite mixture was systematically studied using a differential scanning calorimeter (DSC) in terms of model-fitting Malek approach, integral master plot method and model-free advanced isoconversional method of Vyazovkin. The results showed that the thermal decomposition occurred in two stages. The first step was the decomposition process of Ag_2O to form porous silver particles as a complex process with the participation of at least two mechanisms, and the second step corresponded to the structural change from porous particles to silver bulk crystals as a single-step process.

Key words: Ag_2O ; advanced isoconversional method; kinetic modeling; Sestak–Berggren model

1 Introduction

Silver (I) oxide is a highly thermally stable silver source. Kinetic modeling of Ag_2O decomposition can provide predictive information on the decomposition control. Firstly, the kinetics and mechanism of the Ag_2O decomposition was studied by L'VOV [1], who drew attention to the autocatalytic characteristic of the reaction. These specific features of the reaction are later confirmed by other authors [2–6]. The temperatures in these studies varied from about 173 to 399 °C, and the activation energies were in the 118–180 kJ/mol intervals. The results showed that the kinetics of the thermal decomposition of Ag_2O occurred using a scheme of dissociative evaporation of Ag_2O in the form of free Ag atoms and O_2 molecules with simultaneous condensation of Ag atoms. Moreover, in the initial stage of the decomposition, which provides condensation of the vapor of the low-volatile product at the interface zone and energy transfer to the reactant, the decomposition precedes much slowly [1–6].

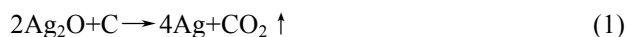
KHAYATI et al [7] investigated the kinetics of the isothermal decomposition of Ag_2O with graphite under argon atmosphere for a non-activated sample and mechanically activated sample. The results showed that the mechanisms of nuclei growth in thermal and mechanochemical reduction are diffusion controlled and

interface controlled, respectively.

To the best of our knowledge, there is not a great deal of information in literature about the kinetics of the thermal or thermochemical decomposition of Ag_2O . In this work, we investigated the kinetics of Ag_2O decomposition using model-fitting Malek approach, integral master plot approach and model-free advanced isoconversional approach of Vyazovkin in a systematic and comparative way. Our current contribution will provide the comprehensive data to better understand the mechanism of the dissociation.

2 Experimental

Starting materials were pure Ag_2O powder (99% purity, 5–40 μm , Merck) and graphite (99.9% purity, 10–50 μm , Merck). Ag_2O with 40% (mole fraction) of extra carbon was reduced according to reaction (1):



To evaluate the kinetics parameters, 30 mg of the mixture of Ag_2O and graphite was selected in each trial. Thermal analyses of Ag_2O decomposition were performed under argon flow using a STA409PG device equipped with thermogravimetric system and an Al_2O_3 crucible. Non-isothermal differential scanning calorimetry (DSC) analyses were carried out from 75 °C to 500 °C at the heating rates of 5, 10, 15 and

20 °C/min. The flow rate of argon was adjusted to 500 mL/min. The samples were characterized by SEM (JEOLJSM 5310) equipped with an energy dispersive spectrometer (EDS) (Oxford Instrument).

3 Theoretical

The rate of decomposition reaction, in general, can be described in terms of two functions, $k(T)$ and $f(\alpha)$, thus,

$$\frac{d\alpha}{dt} = \beta \left(\frac{d\alpha}{dT} \right) = k(T)f(\alpha) \quad (2)$$

where α is the degree of conversion (i.e., proportional to the area under the DSC curve), t is the reaction time, T is the thermodynamics temperature, β is the heating rate, $k(T)$ is the rate constant and $f(\alpha)$ is the function of reaction mechanism. The dependence of the reaction rate constant on temperature can be described by Arrhenius equation:

$$k(T) = A \exp \left(-\frac{E_a}{RT} \right) \quad (3)$$

where A is the pre-exponential factor, R is the gas constant and E_a is the apparent activation energy.

A relatively complete thermal analysis kinetics method proposed by MALEK et al [8,9] was chosen to analyze the non-isothermal experiment data. The Malek method contains two functions, $y(\alpha)$ and $z(\alpha)$, which was used to find the appropriate kinetic model of the studied process. They are as follows [8,10]:

$$y(\alpha) = \left(\frac{d\alpha}{dt} \right) \exp(u) = \beta \left(\frac{d\alpha}{dT} \right) \exp(u) \quad (4)$$

$$z(\alpha) = P(u) \left(\frac{d\alpha}{dt} \right) \left(\frac{T}{\beta} \right) = P(u) \left(\frac{d\alpha}{dT} \right) T \quad (5)$$

where $u = E_a/(RT)$ and $P(u)$ are the expression of the temperature integral, which can be well approximated by the fourth rational expression of MALEK [8] and SENUM and YANG [11] as Eq. (6):

$$P(u) = \frac{u^3 + 18u^2 + 88u + 96}{u^4 + 20u^3 + 120u^2 + 240u + 120} \quad (6)$$

Function $y(\alpha)$ is proportional to function $f(\alpha)$, being characteristic for a given kinetic model. The shapes and maxima of both $y(\alpha)$ and $z(\alpha)$ functions, normalized within the (0, 1) interval, provide valuable information for determination of the most appropriate kinetic model to describe the studied process.

According to the Malek method, E_a must be determined independently using other methods. These methods can be generally categorized as isoconversional and model-fitting methods. Generally, isoconversional methods give accurate values of the activation

energy [12,13].

Model fitting methods depend on the reaction model and also assume the Arrhenius temperature dependent of the rate constant $k(T)$. They do not achieve a clean separation between the temperature dependence $k(T)$ and the reaction model $f(\alpha)$. In addition, the temperature sensitivity of the reaction rate depends on α . To the best of our knowledge, the best known model-fitting methods are the Coats–Redfern (CR) method [14] and the invariant kinetic parameters (IKP) method [15].

Isoconversional methods can be generally divided into integral method and differential method. The integral method relies on the temperature integral approximation, mainly including the Kissinger–Akahira–Sunose (KAS) method [16–19] and the Flynn–Wall–Ozawa (FWO) method. The differential method, such as the Friedman method [10,20], employs the instantaneous values of conversion degree rates, and this method is therefore sensitive to experimental noise. The method uses numerical differentiation of the experimental data to estimate the reaction rates.

There are more complex “model free” methods, like the non-linear isoconversional method by VYAZOVKIN and WIGHT [21] and VYAZOVKIN [22], solutions of which can only be obtained using computer algorithms.

The differential method includes some measurement uncertainties. Because the measurement of the reaction rate $\beta d\alpha/dt$ (or $d\alpha/dt$) is sensitive to the determination of the baseline and calibration of the thermal analysis equipment, the significant inaccuracy in the rate values can be introduced into due to the difficulty of determining the baseline [19]. Also, inaccuracies arise when the reaction heat shows a noticeable dependence on heating rate [23]. It can be pointed out that applying the limited but quantifiable deviation in integral method would be better than the differential method, whose deviations are difficult to calculate [19].

An advanced isoconversional method described by VYAZOVKIN et al [24–27], namely Vyazovkin method, is an effective method to detect the complexity of the reaction and evaluating the activation energy as a function of α . Furthermore, this method has a great calculation accuracy and versatile applicability for various heating programs [23]. Specifically, the Vyazovkin method is applicable to a non-isothermal kinetic process with a series of linear heating, which can be written as [28]

$$\Phi(E_a) = \sum_{i=1}^n \sum_{j \neq i}^n \frac{I(E_a, T_{\alpha,i}) \beta_j}{I(E_a, T_{\alpha,j}) \beta_i} \quad (7)$$

$$I(E_a, T_{\alpha,i}) = \int_{T_{\alpha-\Delta\alpha}}^{T_{\alpha}} \exp \left(-\frac{E_a}{RT_{\alpha}} \right) dT \quad (8)$$

where β_i and β_j represent the different heating rates; T_{α}

and $T_{\alpha-\Delta\alpha}$ are the reaction temperatures corresponding to $\alpha-\Delta\alpha$, respectively. Minimizing Eq. (7) for each α with a certain conversion increment (usually $\Delta\alpha=0.02$) results in the correction of E_a with α . The detailed descriptions to use the Vyazovkin method to treat calorimetric data can be acquired elsewhere [27,29].

4 Results and discussion

4.1 Characterization of precursor

Figure 1(a) shows the SEM image of precursor. To determine the composition of Ag_2O and graphite particles, the EDS point chemical analysis has been

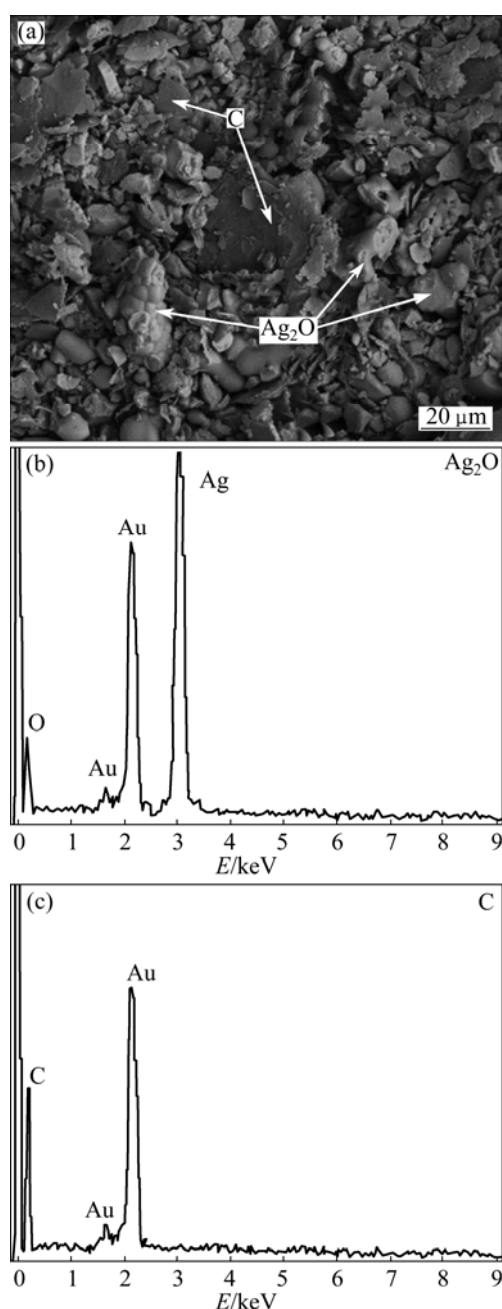


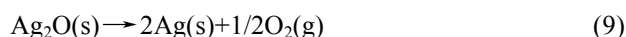
Fig. 1 SEM image (a) and EDS analysis of as-received Ag_2O (b) and graphite (c) mixture

employed with the results shown in Figs. 1(b, c). Accordingly, the graphite particles have been elongated into flake-shaped morphology, while Ag_2O particles show a relatively equiaxed particles morphology with more uniform distribution of particle size. It can be concluded that due to insufficient mixing, the graphite is not activated and prevents clean Ag_2O -to-graphite from contact.

4.2 Thermal decomposition of Ag_2O powder

The DSC curves of thermal decomposition of Ag_2O at four different heating rates (5, 10, 15 and 20 °C/min) are shown in Fig. 2. Accordingly, due to insufficient mixing, the graphite was not activated and thermal decomposition of Ag_2O occurred as an endothermic reaction. As shown, the thermal decomposition of Ag_2O occurred in two steps.

Firstly, occurring at 197–217 °C can be attributed to the Ag_2O decomposition to porous silver particles as follows [30–32]:



The shoulder at the end of the first peak (for all heating rates) probably can be attributed to the formation of aggregates of silver "atoms" from solid solution in Ag_2O crystal as a separate phase, i.e., Ag atoms [6]. On the other hand, each peak or shoulder represents at least one reaction step. So, from Fig. 2, it is obvious that the first step of thermal decomposition of Ag_2O is a complex process with the participation of at least two different mechanisms.

Secondly, occurring at 362–400 °C can be attributed to a structure change from porous particles to silver bulk crystals [33].

Figure 2 shows that, as the heating rate is increased, the endothermic area is shifted to a higher temperature range with expanded peak width. Also, onset reaction temperatures, peak temperatures, and end temperatures (Table 1) are enhanced with increasing heating rate.

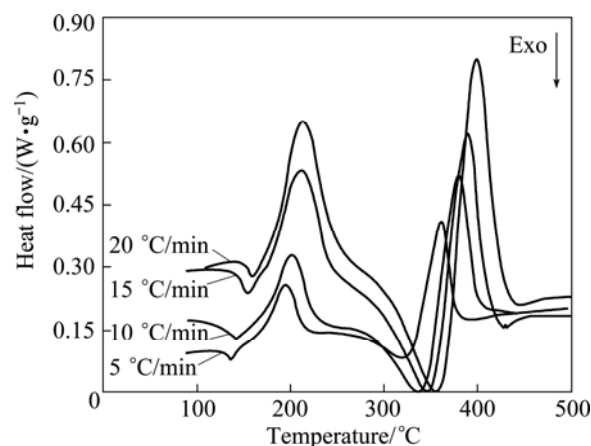


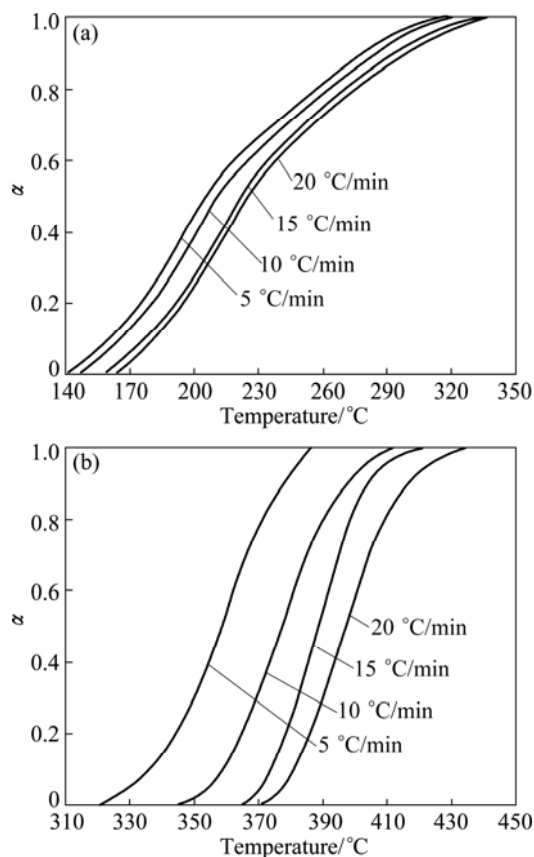
Fig. 2 DSC heat flow rate as function of temperature

Table 1 Typical parameters of Ag₂O thermal decomposition

Step No.	$\beta/$ (°C·min ⁻¹)	$T_{\text{onset}}/$ °C	$T_{\text{peak}}/$ °C	$T_{\text{end}}/$ °C	Reaction heat/(J·g ⁻¹)
The first step	5	140.157	197.638	319.685	52.727
	10	146.457	206.997	333.858	58.722
	15	158.268	213.011	344.094	103.343
	20	163.78	216.887	352.756	124.091
The second step	5	322.331	362.522	386.55	30.959
	10	342.105	381.436	415.479	34.774
	15	363.012	390.105	425.817	47.692
	20	367.397	399.562	435.968	65.806

Integration of the peak area produces the reaction enthalpies at different heating rates (Table 1). The reaction enthalpies are hardly affected by the heating rate, and they change within a range of 52.727–124.091 J/g and 30.959–66.806 J/g for the first and second step, respectively.

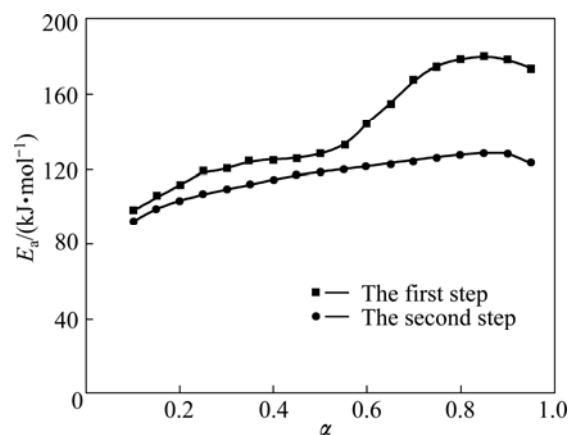
The conversion curves (α - T curves) for non-isothermal thermal decomposition of Ag₂O at different heating rates and for each step of Ag₂O decomposition are indicated in Fig. 3. These conversional curves exhibit the sigmoid profile, and the curves shift towards the higher temperature as the heating rate increases, namely,

**Fig. 3** Changes of α as function of temperature for the first (a) and second (b) steps of Ag₂O thermal decomposition

the higher the heating rate, the higher the temperature for the reaction to reach the identical conversion.

4.3 Calculation of activation energy (E_a)

The dependence of E_a versus α for each step of thermal decomposition of Ag₂O is presented in Fig. 4.

**Fig. 4** Dependence of E_a on α

It can be pointed out that if E_a is independent of α , the decomposition process is dominated by a single reaction step [34,35]; on the contrary, a significant variation of E_a with α should be interpreted in terms of multi-step mechanism [34–37]. Therefore, from Fig. 4, it is obvious that the first step of Ag₂O decomposition (i.e. Ag₂O decomposition to porous silver particles) is a complex process with the participation of at least two mechanisms.

In addition, it can be considered that the E_a is independent of α if the difference between the maximum and minimum values of E_a is less than 20%–30% of the average E_a [28,38]. Therefore, the dependence of E_a on α , according to Vyazovkin method, for the first step can be divided into two different regions ($\alpha \leq 0.55$ and $\alpha > 0.55$). In the first region ($\alpha \leq 0.55$), in the range of $0.10 \leq \alpha < 0.25$, we have a linear increase of E_a with α , while in the range of $0.25 \leq \alpha \leq 0.55$, E_a presents an almost stable behavior. In the second region ($\alpha > 0.55$), in the range of $0.55 < \alpha < 0.70$, we have a linear increase of E_a with α , while in the range of $0.75 \leq \alpha \leq 0.95$, E_a presents an almost stable behavior. The first and second regions of the first step of Ag₂O decomposition can be attributed to the decomposition of Ag₂O decomposition to porous silver particles and the formation of aggregates of silver atoms, respectively [6]. Based on the obtained results from the Vyazovkin method, the average values of E_a for the first and second regions of the first step of Ag₂O decomposition are 118.953 and 168.871 kJ/mol, respectively (Fig. 4).

In the second step of Ag₂O decomposition, in the range of $0.10 \leq \alpha \leq 0.90$, we have an almost linear increase of E_a with α . In addition, from Fig. 4, it is obvious that

the difference between the maximum and minimum values of E_a is less than 20%–30% of the average E_a (118.933 kJ/mol). Therefore, E_a presents an almost stable behavior and the second step of Ag_2O decomposition is a single-step process.

Basically, the Kissinger method is applied to analyzing the DSC data because it is independent of any pre-assumption about the thermal decomposition mechanism [17]. In this method, the activation energy is calculated from straight line fitting of $\ln(\beta/T_p^2)$ versus $1000/T_p$ plots according to Eq. (10) [17]:

$$\frac{d[\ln(\beta/T_p^2)]}{d(1/T_p)} = -E_a/R \quad (10)$$

where T_p is the temperature corresponding to the maximum of thermal decomposition rate. Figure 5 shows the results of the Kissinger method for the first step of Ag_2O decomposition. It is obvious, however, that the Kissinger method accuracy is very high, but the first step of Ag_2O decomposition is a complex process with the participation of at least two mechanisms.

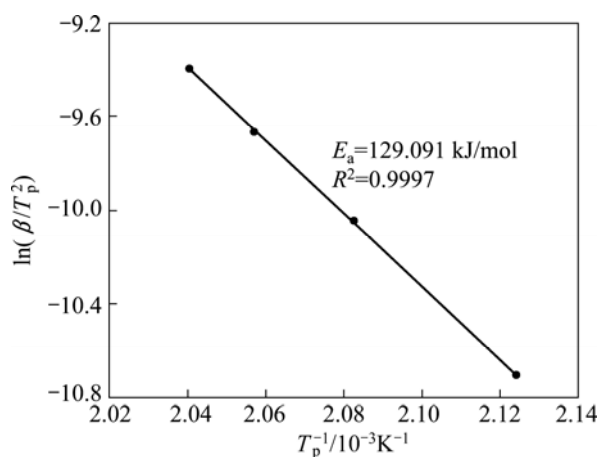


Fig. 5 Kissinger plot for the first step of Ag_2O thermal decomposition

As a result, it is worth to note that the Kissinger method might be the simplest method for analyzing the decomposition process, but produces a single value of the activation energy for any process regardless of its actual kinetic complexity [28]. So, this method is applicable only to the single-step process. Therefore, the Vyazovkin method was applied in this work.

4.4 Characterization of reduction products

Figure 6 shows the typical SEM images and EDS point chemical analysis of precursor after heating at the rate of $15\text{ }^\circ\text{C/min}$ and $\alpha=0.65$. Accordingly, the Ag products are generally round in shape and somewhat porous structure. As shown in Fig. 6, each Ag particle contains many fine crystallites, namely, during the first step Ag_2O is decomposed, and aggregated to each other

in the second stage during further heating. Moreover, EDS spectrum reveals that the agglomerates are pure Ag and no impurities such as carbon and oxygen are detected.

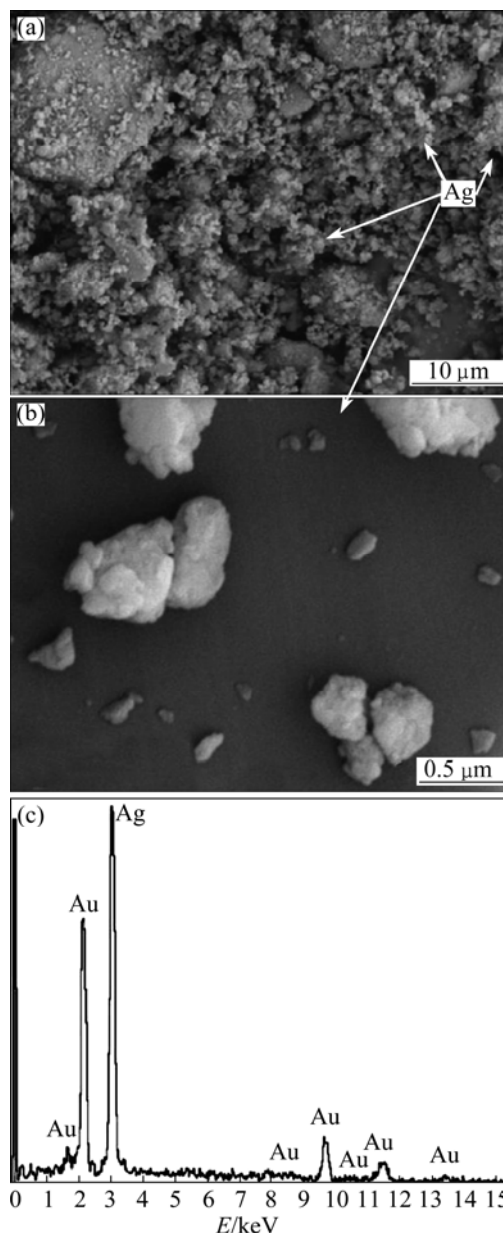


Fig. 6 Typical SEM images (a, b) and EDS analysis (c) of Ag_2O -graphite after heating at rate of $15\text{ }^\circ\text{C/min}$ and $\alpha=0.65$

4.5 Determination of most probable reaction mechanism

Figure 7 shows the changes of reaction rates $\beta(da/dT)$ as a function of conversion degree for each step of Ag_2O decomposition. $\beta(da/dT)$ increases with the heating rate, and the peak reaction rate appears in the conversion range of 0.383–0.393. This observation likely implies that the heating rate greatly affects the reaction kinetics, but hardly affects the basic reaction mechanisms.

The normalized function curves of $y(\alpha)$ and $z(\alpha)$ are constructed by following Eqs. (4) and (5) for the first step, in the range of $0.1 \leq \alpha \leq 0.55$, and the second step of Ag_2O decomposition (see Figs. 8 and 9).

These figures clearly indicate that the $y(\alpha)$ curves exhibit peak values appearing at conversion, α_M , 0.100 and 0.277–0.337 for the first and second steps, respectively. The $z(\alpha)$ curves show a “C” shaped contour

experiencing a practically isoconversional peak value conversion, α_p^∞ , within 0.399–0.416 and 0.516–0.586 for the first and second steps, respectively. Also, it is independent of the applied heating rate. These findings demonstrate that the effect of heating rate on the whole reaction kinetics can be removed by applying the Malek method to the non-isothermal kinetic data of Ag_2O decomposition.

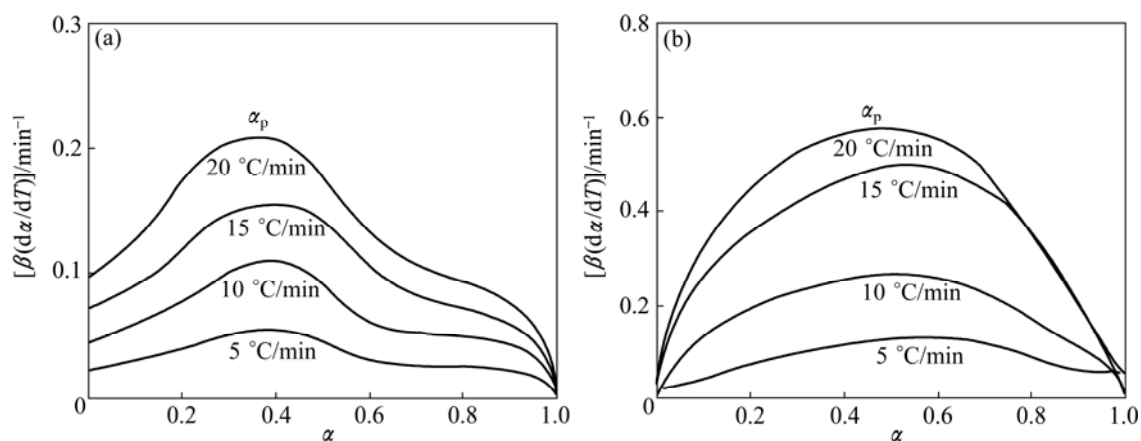


Fig. 7 Reaction rate $\beta(da/dT)$ as function of conversion for the first (a) and second (b) steps of Ag_2O thermal decomposition

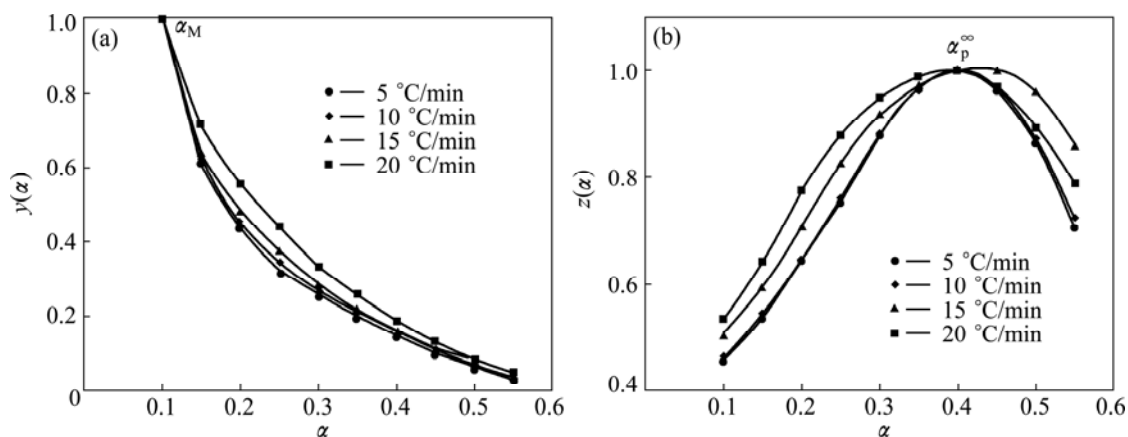


Fig. 8 Plots of normalized $y(\alpha)$ and $z(\alpha)$ against α for the first step of Ag_2O thermal decomposition

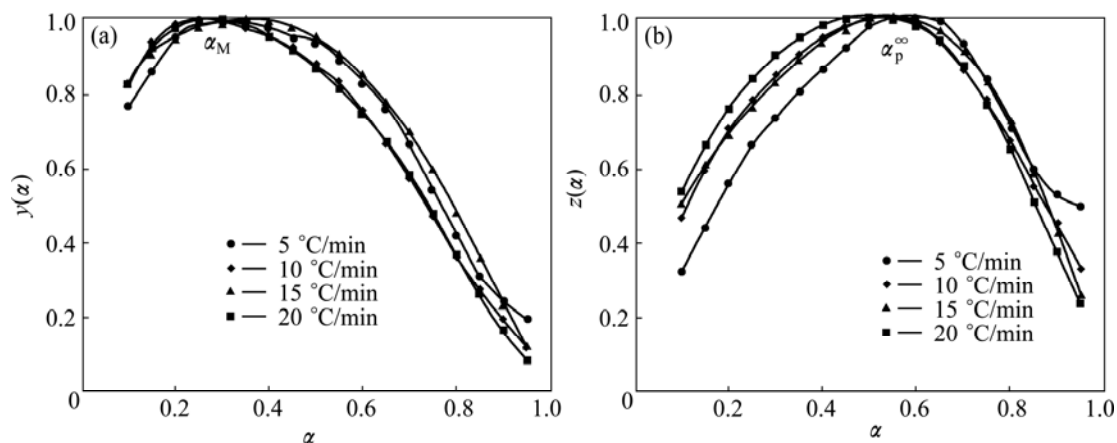


Fig. 9 Plots of normalized $y(\alpha)$ and $z(\alpha)$ against α for the second step of Ag_2O thermal decomposition

The peak value of $y(\alpha)$ appears while conversion is much higher than 0, implying that the reaction is auto-catalyzed [39]. Besides, α_M being larger than zero demonstrates the autocatalytic property of reactions [39]. According to the shape of $y(\alpha)$ and characteristic values of $y(\alpha)$ and $z(\alpha)$, we may draw a conclusion that the truncated Sestak–Berggren model (SB(m , n) model) [8,40–42] or extended Prout–Tompkins model (the regular Prout–Tompkins model is $f(\alpha)=\alpha(1-\alpha)$) is suitable for kinetic modeling of Ag_2O decomposition. In addition, from Table 2, one can observe that $0 < \alpha_M < \alpha_p^\infty$ and $\alpha_p^\infty \neq 0.632$, which strongly indicates that the truncated SB(m , n) [41,42] is suitable. It is worth to note that the truncated SB(m , n) is not suitable for kinetic modeling of the first step of the Ag_2O decomposition in the range of $0.55 < \alpha \leq 0.95$ because the SB(m , n) conditions are not indefeasible.

Table 2 Characteristic peak conversion values of α_p , α_M and α_p^∞

Step No.	$\beta/(\text{°C} \cdot \text{min}^{-1})$	α_p	α_M	α_p^∞
The first step	5	0.393	0.100	0.403
	10	0.392	0.100	0.402
	15	0.391	0.100	0.416
	20	0.383	0.100	0.399
The second step	5	0.581	0.314	0.586
	10	0.552	0.275	0.552
	15	0.556	0.337	0.559
	20	0.516	0.277	0.516

SESTAK and BERGGREN [42] have proposed an empirical model for $f(\alpha)$ as follows:

$$f(\alpha) = \alpha^m (1-\alpha)^n [-\ln(1-\alpha)]^s \quad (11)$$

According to the Sestak–Berggren model, the combination of m , n , and s can represent a number of different reaction models. The increasing value in the exponent m indicates a more important role of the precipitated phase on the overall kinetics. It also presents that a higher value of the second exponent ($n > 1$) indicates increasing reaction complexity [41,42]. Furthermore, value of the exponent s can give us information about the nucleation properties of a given complex mechanism of the tested process [41,42]. It is generally used in truncated form ($s=0$ in Eq. (11)). The truncated Sestak–Berggren model is an example of an autocatalytic model. According to the truncated SB(m , n) model and Eq. (3), Eq. (2) can be transformed into [41,42]

$$\frac{d\alpha}{dt} = \beta \left(\frac{d\alpha}{dT} \right) = A \exp\left(-\frac{E_a}{RT}\right) \alpha^m (1-\alpha)^n \quad (12)$$

where m and n are the reaction orders, and other

parameters have the same meaning as the above equations. The ratio of m to n , p , equals $\alpha_M/(1-\alpha_M)$ according to the results by MALEK [8]. Thus, Eq. (12) can be transformed into

$$\ln \left[\beta \left(\frac{d\alpha}{dT} \right) \exp\left(\frac{E_a}{RT}\right) \right] = \ln A + n \ln[\alpha^p (1-\alpha)] \quad (13)$$

The kinetic parameter n can be derived from the slope of $\ln[\beta(d\alpha/dT)\exp(E_a/RT)]$ versus $\ln[\alpha^p(1-\alpha)]$ for $0.20 \leq \alpha \leq 0.95$ [43], and the intercept is $\ln A$, and $m=pn$. These data perform good linearity (Fig. 10). The values of m , n and $\ln A$ can be calculated from the intercept and slope of these fit straight lines, which are listed in Table 3. The higher value of the kinetic parameter n ($n > 1$) indicates the increasing complexity of the decomposition process [42]. Therefore, the first step of the Ag_2O decomposition is a process with the complexity higher than that of the second step.

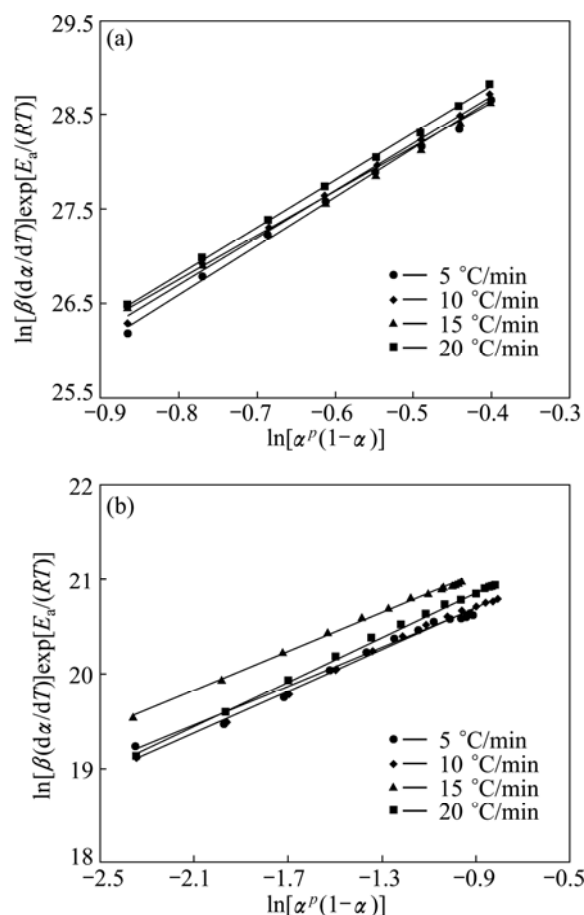


Fig. 10 Plots of $\ln[\beta(d\alpha/dT)\exp(E_a/RT)]$ vs $\ln[\alpha^p(1-\alpha)]$ at different heating rates for the first (a) and the second (b) steps of Ag_2O thermal decomposition

By substituting calculated kinetic parameters (n , m and $\ln A$ in Table 3 along with previously calculated E_a) into Eq. (13), the explicit rate equations for the first step, in the range of $0.10 \leq \alpha \leq 0.55$, and the second step

Table 3 Calculated kinetic parameters m , n and $\ln A$ for SB(m , n) model

Step No.	$\beta/(^{\circ}\text{C}\cdot\text{min}^{-1})$	m	n	$\ln A$
The first step	5	0.570	5.132	30.700
	10	0.556	5.004	30.700
	15	0.516	4.647	30.468
	20	0.550	4.950	30.766
The second step	5	0.473	1.034	21.610
	10	0.421	1.111	21.710
	15	0.517	1.019	21.969
	20	0.450	1.171	21.908

of the Ag_2O thermal decomposition are obtained by Eqs. (14) and (15), respectively.

$$\frac{d\alpha}{dt} = 2.065 \times 10^{13} \exp\left(-\frac{118953}{RT}\right) \alpha^{0.548} (1-\alpha)^{4.933},$$

$$\alpha \in [0.10, 0.55] \quad (14)$$

$$\frac{d\alpha}{dt} = 2.933 \times 10^9 \exp\left(-\frac{118933}{RT}\right) \alpha^{0.465} (1-\alpha)^{1.084},$$

$$\alpha \in [0.10, 0.95] \quad (15)$$

To check the predictability of the explicit kinetic models (Eqs. (14) and (15)), the predicted rate curves of different heating rates are calculated (Fig. 11). From

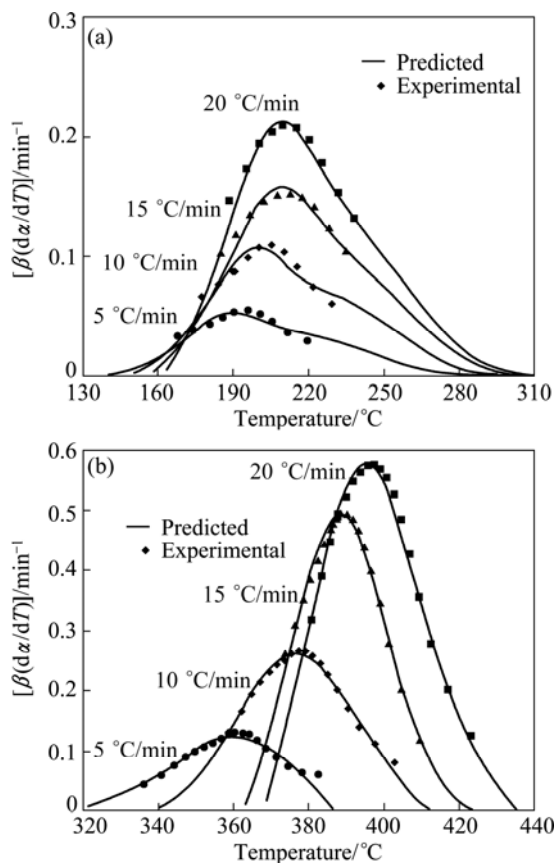
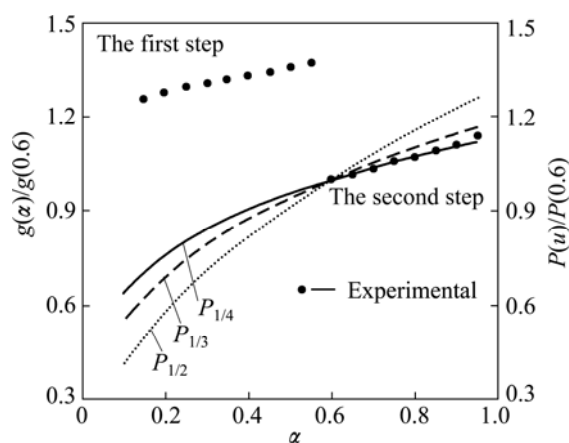
**Fig. 11** Comparison of experimental reaction rate and that predicted from SB(m , n) model vs temperature for the first (a) and second (b) steps of Ag_2O thermal decomposition

Fig. 10, there is a very good agreement between the experimentally obtained $\beta(d\alpha/dT)-T$ curves and numerically calculated rate curves for the Ag_2O decomposition, using the corresponding kinetic triplets. Hence, the truncated SB(m , n) model is suitable for predicting kinetic rate for describing the progress of the non-isothermal decomposition of Ag_2O powder.

To determine the suitable kinetic model for the first step of Ag_2O decomposition ($0.55 < \alpha \leq 0.95$), the integral master plot method can be applied. Figure 12 shows the theoretical integral master curves for $P_{1/4}(f(\alpha)=4\alpha^{3/4})$, $P_{1/3}(f(\alpha)=3\alpha^{2/3})$ and $P_{1/2}(f(\alpha)=2\alpha^{1/2})$ kinetic models, and also the experimental master curves (symbols). It can be seen from Fig. 12 that the experimental master curves (at all heating rates) are in good agreement with the theoretical master curve for $P_{1/4}$ model. As a result, the first step of the Ag_2O thermal decomposition process in the range of $0.10 \leq \alpha \leq 0.55$ follows an autocatalytic model (Eq. (14)), and follows an acceleratory rate model (nucleation mechanism) in the range of $0.55 < \alpha \leq 0.95$.

**Fig. 12** Comparison between theoretical integral master curve for $P_{1/4}$, $P_{1/3}$ and $P_{1/2}$ kinetic models and experimental master curve for the first step of Ag_2O thermal decomposition in range of $0.55 < \alpha \leq 0.95$

5 Conclusions

1) The non-isothermal decomposition kinetics and behaviors of the Ag_2O -graphite powder were systematically studied by DSC technique. The DSC results show that due to insufficient mixing of graphite with Ag_2O , graphite is not activated and participated in the reaction. Consequently, the thermal decomposition of Ag_2O occurs as endothermic reaction in two steps. Firstly, occurring at 197–217 $^{\circ}\text{C}$ can be attributed to the decomposition of Ag_2O to porous silver particles. The shoulder at the end of the first peak probably can be attributed to the formation of silver atoms aggregates in an oxide lattice saturated with silver. Secondly, occurring at 362–400 $^{\circ}\text{C}$ can be attributed to a structure change

from porous Ag particles to silver bulk crystals.

2) In addition, the model-free kinetic analysis with the Vyazovkin method revealed the relationship between E_a and α , which could reflect the kinetic complexity of the Ag_2O thermal decomposition. The results show that the first step of Ag_2O decomposition is a complex process with the participation of at least two mechanisms and the second step of Ag_2O decomposition is a single-step process.

3) Based on the results obtained by Vyazovkin method calculation, the average values of E_a for the first ($0.10 \leq \alpha \leq 0.55$) and the second ($0.55 < \alpha \leq 0.95$) regions of the first and the second steps of Ag_2O decomposition are 118.953, 168.871 kJ/mol and 118.933 kJ/mol, respectively.

4) The model-fitting process kinetic analysis with the Malek method disclosed the typical autocatalytic characteristics for the first ($0.10 \leq \alpha \leq 0.55$) and the second step of Ag_2O decomposition. The further analysis confirms that the Sestak–Berggren model is suitable for kinetic prediction of the process rates, and a complete set of the kinetic parameters for each model is calculated. Besides, the explicit rate equations are established. The kinetic predictions of process rates from the rate equations are made in good agreement with the experimental results. Furthermore, the first step of the Ag_2O thermal decomposition process in the range of $0.10 \leq \alpha \leq 0.55$ follows an autocatalytic model, and follows an acceleratory rate model (nucleation mechanism) in the range of $0.55 < \alpha \leq 0.95$.

5) In the future, we plan to analyze the effect of mechanical activation of starting materials by high energy planetary ball mill. We also try to investigate the effect of practical parameter of milling such as milling time on kinetics consideration of thermochemical decomposition of Ag_2O powder.

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Ag_2O –石墨混合物非等温分解的动力学分析

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摘要: 根据模型拟合 Malek 法、积分主曲线法和无模型的先进等转化率法, 采用差示扫描量热仪系统研究了 Ag_2O –石墨混合物热分解的动力学。结果表明, 热分解包含两个阶段, 在第一阶段, Ag_2O 分解形成多孔银颗粒, 该过程复杂, 至少存在两种机制; 在第二阶段, 多孔粒子发生结构变化而形成大块银晶体, 此过程为单步过程。

关键词: Ag_2O ; 先进的等转化率法; 动力学模型; Sestak–Berggren 模型

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